

Sphere to cylinder transition in a single phase microemulsion system: A theoretical investigation

Edgar M. Blokhuis

Colloid and Interface Science, Leiden Institute of Chemistry, Gorlaeus Laboratories, P.O. Box 9502, 2300 RA Leiden, The Netherlands

Wiebke F. C. Sager

KFA Jülich GmbH, Forschungszentrum, IFF, D-52425 Jülich, Germany

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The sphere to cylinder transition in a one-phase droplet microemulsion system is studied theoretically. Within the framework of the curvature energy model by Helfrich, it was already shown by Safran *et al.* [*J. Phys. (France) Lett.* **45**, L-69 (1984)] that for a certain range of the curvature parameters (rigidity constants and spontaneous curvature), a transition occurs from spherical droplets to infinitely long cylinders through a region where both spheres and cylinders are present. Our aim is to further investigate this region in a *quantitative* way by including—in addition to curvature energy—translation entropy, cylinder length polydispersity, and radial polydispersity. In this way we are able to obtain structural information on the spheres and cylinders formed, their respective volume fractions, and polydispersity, and provide a more detailed comparison with experimental results. © 2001 American Institute of Physics. [DOI: 10.1063/1.1380428]

I. INTRODUCTION

Over the past three decades it has been shown that microemulsions are structurally well-defined self-organizing mixtures of water, oil, and surfactants that can form a wide variety of thermodynamically stable phases. These comprise phases consisting of (more or less spherical) droplets of water in oil (w/o microemulsions or L_2 -phase) or oil droplets in water (o/w microemulsion or L_1 -phase), as well as bicontinuous mono- and bilayer phases. If the temperature and/or ionic strength of the aqueous phase is varied, a rich phase behavior is generally revealed,^{1,2} whereby microemulsion phases can coexist with water and/or oil excess phases as well as liquid crystalline phases forming two- and three-phase equilibria.

Ample experimental evidence from, e.g., electric birefringence (Kerr-effect), dielectric spectroscopy, fluorescence quenching, turbidity, and temperature jump experiments, has been presented for droplet aggregation in the L_1 and L_2 -phases.³ Furthermore, a considerable jump (2–3 orders of magnitude) in conductivity has been reported for w/o AOT microemulsions when the temperature or amount of internal phases is increased.⁴ It is still a matter of debate whether this increase in conductivity is a result of charge transfer between aggregated droplets or whether the aggregated droplets open up forming interconnected cylindrical structures. Even though the formation of cylindrical structures is extensively documented experimentally in micellar systems,^{5,6} their existence is less well-established in microemulsions.

Recent evidence for the formation of cylindrical structures has been supplied by SAXS and SANS scattering studies in the L_1 and L_2 -phases. The usual interpretation of the scattering data in this region focuses on a fit of (aggregating) spherical droplets characterized by an average radius, radial polydispersity, and stickiness parameter describing the de-

gree of aggregation.^{7,8} It was found by Ilgenfritz *et al.*⁹ and Glatter *et al.*¹⁰ that also cylindrical structures start to become present. Only recently could a more *quantitative* study be undertaken of this structural transition from spherical droplets to cylindrical structures in the one-phase microemulsion system.¹¹ Using SAXS, the structural changes in w/o AOT microemulsions were investigated as a function of temperature (15–60 °C), salt concentration (up to 0.6% NaCl), water/AOT molar ratio (25–60), and droplet weight fraction (2%–20%).¹¹

Theoretically it now seems well-established that the Helfrich free energy¹² describing the curvature free energy of the interfacial surfactant layer can help understand the global features of the microemulsion phase diagram.^{13,14} As in the experimental situation, while a lot is understood when it concerns the description of *spherical* microemulsion droplets, relatively little is known on the formation of *cylindrical* structures. Pivotal work has been done by Safran and co-workers.^{15–17} In 1984 the phase diagram of the sphere to cylinder transition was published based on Helfrich's curvature energy but neglecting entropy effects,^{14,15} while more recently entropy effects were included to predict closed-loop coexistence regions in a system consisting of spheres and cylinders.¹⁶ It is our intention in this article to extend the work by Safran and co-workers and to study the sphere to cylindrical transition in the one-phase region in more detail. Our aim is to investigate this in a *quantitative* way by including translation entropy, cylinder length polydispersity, and radial polydispersity.

Our theory has three important ingredients:

(1) Curvature free energy: The Helfrich form of the free energy is used to describe the interfacial free energy of the curved surfactant monolayer that separates the oil and the water phase,¹²

$$F_{\text{curv}} = \int dA \left[-\frac{2k}{R_0} J + \frac{k}{2} J^2 + \bar{k} K \right]. \quad (1.1)$$

The above free energy features an integral over the whole surface area, A , of the total curvature, $J = 1/R_1 + 1/R_2$ and Gaussian curvature, $K = 1/(R_1 R_2)$ with R_1 and R_2 the principal radii of curvature at a certain point on the surface A . Three curvature parameters are introduced: $1/R_0$, the inverse radius of spontaneous curvature, k the rigidity constant of bending, and \bar{k} the rigidity constant associated with Gaussian curvature. In the following analysis, we treat these curvature parameters as unknown and construct our phase diagrams in terms of them. However, we do know from experiments that k and \bar{k} are approximately constant over the temperature range considered and of the order of 1 or a few $k_B T$,¹⁸ while the inverse radius of spontaneous curvature changes significantly as a function of temperature (approximately linear) and can even change sign at the so-called inversion temperature \bar{T} .¹⁹ Therefore, in the comparison with experiments we consider k and \bar{k} as constants and treat $1/R_0$ as our “temperature variable,” $1/R_0 \propto (T - \bar{T})$.

(2) Entropy: Although the consideration of the curvature free energy alone already gives good qualitative insight into the microemulsion phase diagram, entropy needs to be considered in any more quantitative analysis. Entropy is generally responsible for the occurrence of polydispersity, which is an important feature of microemulsion systems, and it will smoothen structural transitions in the one-phase region like the sphere to cylindrical transition that we consider here. The theory for including entropy in microemulsion systems is however not free of controversy in the literature.²⁰ In this article we investigate a number of different expressions for the entropic contribution to the free energy to find out which aspects are model dependent and which aspects are more generally valid.

(3) Constraints: Two constraints have to be considered. First, the total volume, V_{tot} , inside the spheres and cylinders is determined by the amount of internal phase present, for instance the amount of water when we consider water-in-oil microemulsions. Second, the total surface area, A , is determined by the amount of surfactant in the system. When we minimize the free energy consisting of the curvature free energy and the entropic contribution to the free energy, these two constraints have to be taken into account. One way to take these constraints into account is to add Lagrange multipliers (which we will call σ and $-\Delta p$) to the free energy.

We start, in Sec. II, with a reinspection of the phase diagram of the sphere to cylindrical transition in which only the curvature free energy is taken into account with the above constraints neglecting the contribution of entropy. This phase diagram was first published by Safran^{14,15} and it already shows many features of the phase diagrams calculated in later sections when entropy is taken into account. In Sec. III, translation entropy is included and the cylinder length polydispersity is considered, while in Sec. IV also the polydispersity in the radius of the sphere and the cylinder is taken into account. In the final section we summarize our findings and discuss the limitations of the theory presented.

II. NO ENTROPY: SPHERES AND INFINITELY LONG CYLINDERS

The curvature free energy of N_s spheres with radius R_s is derived by inserting $J = 2/R_s$ and $K = 1/R_s^2$ into Eq. (1.1),

$$\frac{F_{\text{curv},s}}{\pi k} = N_s \left\{ -\frac{16}{R_0} R_s + 4(2+x) \right\}, \quad (2.1)$$

with $x \equiv \bar{k}/k$ defined as the ratio of the two rigidity constants. The total volume and surface area are given by

$$A = N_s 4 \pi R_s^2, \quad (2.2)$$

$$V_{\text{tot}} = N_s \frac{4}{3} \pi R_s^3.$$

The curvature free energy of N_c cylinders with radius R_c and length $L \gg R_c$ (so that we can neglect the curvature energy of the ends of the cylinder) is derived by inserting $J = 1/R_c$ and $K = 0$ into Eq. (1.1),

$$\frac{F_{\text{curv},c}}{\pi k} = N_c L \left\{ -\frac{4}{R_0} + \frac{1}{R_c} \right\}. \quad (2.3)$$

The volume and surface area are given by

$$A = N_c 2 \pi R_c L, \quad (2.4)$$

$$V_{\text{tot}} = N_c \pi R_c^2 L.$$

The free energy of a system containing both spheres and cylinders is the sum of the above free energies of N_s spheres and N_c cylinders. Instead of N_s and N_c as parameters, it is more convenient to use the *volume fractions* of spheres and cylinders, v_s and v_c , defined as

$$v_s \equiv \frac{N_s}{V_{\text{tot}}} \frac{4}{3} \pi R_s^3, \quad (2.5)$$

$$v_c \equiv \frac{N_c}{V_{\text{tot}}} \pi R_c^2 L.$$

The total free energy then becomes

$$\frac{F_{\text{curv}}}{k V_{\text{tot}}} = v_s \left\{ -\frac{12}{R_0} \frac{1}{R_s^2} + 3(2+x) \frac{1}{R_s^3} \right\} + v_c \left\{ -\frac{4}{R_0} \frac{1}{R_c^2} + \frac{1}{R_c^3} \right\}, \quad (2.6)$$

with the volume and area constraints written as

$$v_s + v_c = 1, \quad (2.7)$$

$$\frac{3v_s}{R_s} + \frac{2v_c}{R_c} = \frac{A}{V_{\text{tot}}} \equiv \frac{1}{\omega}.$$

The ratio between the total volume and surface area defines the length scale ω . In general, ω depends on the size of the surfactant molecules and the molecules constituting the internal phase. Specifically, $A = n_{\text{surf}} a_{\text{surf}}$, with n_{surf} the number of surfactant molecules and a_{surf} the surface area taken in by a surfactant molecule; $V_{\text{tot}} = n_{\text{int}} v_{\text{int}}$, with n_{int} the number of molecules in the internal phase and v_{int} the volume per molecule taken in by the internal phase.

In the following we express all lengths (R_s , R_c , R_0 , and L) in terms of ω . For instance, in the case that the minimi-

zation of the total curvature free energy yields $v_s=1$ (only spheres), one immediately finds from the second constraint in Eq. (2.7) that $R_s=3$, in units of ω . Analogously, when only cylinders are present ($v_c=1$), the radius of the cylinder is given by $R_c=2$, in units of ω . In these cases, with only one species present, the radii therefore follow *directly* from the constraints. The free energies are then simply obtained from Eq. (2.6),

$$\begin{aligned}\frac{\omega^3 F_{\text{curv}}}{k V_{\text{tot}}} &= -\frac{4}{3} \frac{1}{R_0} + \frac{1}{9}(2+x), & \text{only spheres} \\ \frac{\omega^3 F_{\text{curv}}}{k V_{\text{tot}}} &= -\frac{1}{R_0} + \frac{1}{8}, & \text{only cylinders.}\end{aligned}\quad (2.8)$$

When we consider the free energy of the system consisting of both spheres and cylinders, the radii have to be determined from the minimization of the free energy in Eq. (2.6) with the constraints in Eq. (2.7). A convenient way is to first solve the volume fractions v_s and v_c in terms of R_c and R_s from Eq. (2.7), yielding

$$v_s = \frac{R_s(2-R_c)}{2R_s-3R_c}, \quad (2.9)$$

$$v_c = \frac{R_c(R_s-3)}{2R_s-3R_c},$$

and insert the result into the free energy in Eq. (2.6). The minimizing equations $\partial F/\partial R_s=0$ and $\partial F/\partial R_c=0$ then yield the following pair of algebraic equations to determine R_s and R_c in terms of x and $1/R_0$:

$$\begin{aligned}\frac{4}{R_0} R_s R_c (R_s^2 + 3R_c^2 - 4R_s R_c) \\ + 6(2+x) R_c^2 (R_s - R_c) - R_s^3 = 0, \\ -\frac{8}{R_0} R_s R_c (R_s^2 + 3R_c^2 - 3R_s R_c) \\ + 6(2+x) R_c^3 + R_s^2 (4R_s - 9R_c) = 0.\end{aligned}\quad (2.10)$$

With the free energies of the three systems (spheres, cylinders, and spheres+cylinders) determined, one is then able to construct the phase boundaries of the transitions between spheres and cylinders and between these two phases and the phase consisting of both spheres and cylinders. It should be realized that the phase consisting of spheres and cylinders is still a *single* phase and not phase separated.

Before showing the complete phase diagram we first need to discuss the transition to two other phases. It turns out that for small values of $1/R_0$ (large R_0) a transition occurs to the *lamellar phase*, while for large values of $1/R_0$ (small R_0) the internal phase is expelled as an excess phase and phase separation occurs. The boundary at which the latter transition occurs is termed the *solubilization limit* or the *emulsification failure* transition. These two phases are now discussed in more detail.

A. Solubilization limit

At the solubilization limit (SL), the internal phase starts to be present as an excess phase. If the volume of the excess phase is denoted by V_0 and v_0 is defined as $v_0 \equiv V_0/V_{\text{tot}}$, the volume constraint in Eq. (2.7) becomes

$$v_0 + v_s + v_c = 1. \quad (2.11)$$

The amount of internal phase that is expelled as an excess phase, v_0 , is determined by a minimization of the free energy with respect to v_0 . The solubilization limit is therefore determined by the minimization equation $\partial F/\partial v_0=0$, with the condition that $v_0=0$ at the solubilization limit. This procedure gives as solubilization limit for the three systems, spheres, cylinders, and spheres+cylinders, the following relations between $1/R_0$ and x :

$$\begin{aligned}\frac{1}{R_0} &= \frac{1}{6}(2+x), & \text{SL, only spheres,} \\ x &= 0, & \text{SL, spheres+cylinders,} \\ \frac{1}{R_0} &= \frac{1}{4}, & \text{SL, only cylinders.}\end{aligned}\quad (2.12)$$

B. Transition to the lamellar phase

The lamellar phase (L_α) is characterized by planar sheets of surfactant films that carry no curvature so that the corresponding curvature energy is zero ($F_{\text{curv}}=0$). Therefore, when the calculated curvature energy of the spheres and cylinders changes sign and becomes *positive*, the free energy for forming spheres and/or cylinders is higher than the free energy associated with the lamellar phase and a (first order) phase transition occurs. The location of the transition to the lamellar phase is thus determined by inserting $F_{\text{curv}}=0$ into Eq. (2.6). One finds

$$\begin{aligned}\frac{1}{R_0} &= \frac{1}{12}(2+x), & L_\alpha, \text{ only spheres,} \\ \frac{1}{R_0} &= \frac{1}{8}, & L_\alpha, \text{ only cylinders.}\end{aligned}\quad (2.13)$$

The expression for $1/R_0$ at the transition to the lamellar phase of the phase comprising spheres *and* cylinders is somewhat tedious and we will not reproduce it here. It can be derived from solving the set of equations in Eq. (2.10) together with the condition,

$$\begin{aligned}\frac{4}{R_0} R_s R_c (R_s^2 - 3R_s + 6R_c - 3R_c^2) + 3(2+x) R_c^2 (R_c - 2) \\ + R_s^2 (3 - R_s) = 0,\end{aligned}\quad (2.14)$$

for R_s , R_c , and $1/R_0$.

The resulting phase diagram as a function of ω/R_0 and $x=\bar{k}/k$ is shown in Fig. 1. With slightly different axes, it was already published by Safran.^{14,15} The upper region is the 2ϕ -region where the microemulsion coexists with the internal phase, the lower region is the lamellar phase (L_α), and in between is the one-phase microemulsion region, which consists of either spheres, or cylinders, or spheres and cylinders

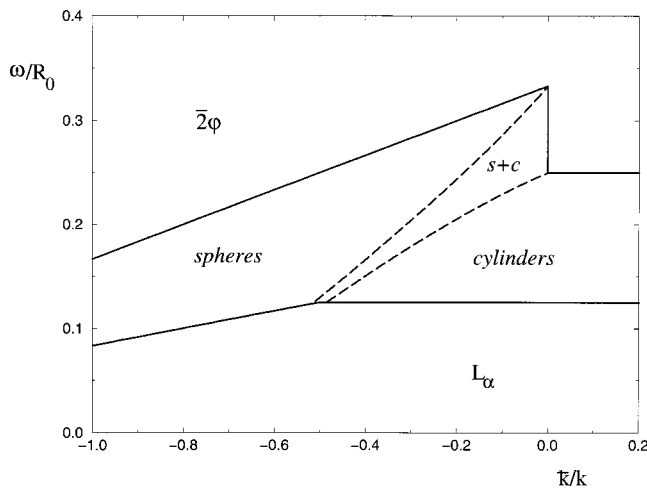


FIG. 1. Microemulsion phase diagram without entropy as a function of ω/R_0 and k/k . The upper region is the $\bar{2}\phi$ -region where the microemulsion coexists with the internal phase. The lower region is the lamellar phase (L_α). In between is the one-phase microemulsion region, which consists of either spheres, or (infinitely long) cylinders, or spheres and cylinders ($s+c$) with sharp transitions (dashed lines) between them.

($s+c$) with sharp transitions (dashed lines) between them. It should be emphasized that in the calculation of the phase diagram in Fig. 1, we have only considered spheres and (infinitely long) cylinders as structures possibly present in the microemulsion phase. More complex structures, such as saddlelike structures, are therefore not considered but are expected to play a role when $x > 0$.

III. TRANSLATIONAL ENTROPY: SPHERES AND SPHEROCYLINDERS

Next, we consider the influence of translational entropy on the phase diagram. For the cylinders, our treatment of the influence of entropy is very much in the spirit of the Flory–Huggins theory of polymers.²² We first consider the entropic contribution to the free energy of spheres only.

A. Spheres

Subdividing the volume V into volume elements $v_{0,s}$ each containing one or none spherical droplets, the entropy contribution of N_s spheres is given by

$$F_{\text{ent},s} = k_B T \left[N_s \ln \left(\frac{N_s v_{0,s}}{V} \right) - N_s \right], \quad (3.1)$$

where k_B is Boltzmann's constant and T the absolute temperature. In writing Eq. (3.1) we have assumed that the total volume occupied by the spherical particles is much smaller than the volume of the vessel, $V_{\text{tot}} \ll V$. The total free energy is the sum of the curvature energy and translational entropy

$$\frac{F_s}{\pi k} = N_s \left\{ -\frac{16}{R_0} R_s + 4(2+x) + t \left[\ln \left(\frac{N_s v_{0,s}}{V} \right) - 1 \right] \right\}, \quad (3.2)$$

where we have defined the reduced temperature, or reduced inverse rigidity constant of bending for that matter, $t \equiv k_B T / (\pi k)$. The volume and area constraints are still given by Eq. (2.2),

$$A = N_s 4 \pi R_s^2, \quad (3.3)$$

$$V_{\text{tot}} = N_s \frac{4}{3} \pi R_s^3.$$

The unknown volume element $v_{0,s}$ plays the role of the de Broglie volume Λ^3 . The cubic root of $v_{0,s}$ is the typical length scale over which a microemulsion droplet needs to be displaced in order for it to “count” as constituting a different state. Its magnitude and scaling with, e.g., the droplet size is a matter of some debate.²⁰ Within the context of statistical mechanical treatments using the curvature energy model,²³ it now seems well-established that the hypothesis of Safran and co-workers^{14,16} to assume $v_{0,s}^{1/3}$ to be of the order of the droplet radius itself is a “reasonable approximation.”²⁰ In the present treatment we therefore take $v_{0,s}$ to be of the order of the droplet size,

$$v_{0,s} \approx \frac{4\pi}{3} R_s^3. \quad (3.4)$$

Introducing the total volume fraction $\phi \equiv V_{\text{tot}}/V$, the free energy can then be written as

$$\frac{F_s}{\pi k} = N_s \left\{ -\frac{16}{R_0} R_s + 4(2+x) + t [\ln(\phi) - 1] \right\}. \quad (3.5)$$

As in Sec. II, when only spheres are present, the spherical droplet radius follows directly from the constraints: $R_s = 3$ (in units of ω). The free energy F_s is then simply obtained by substituting $R_s = 3$ into the expression above. Different expressions for the entropic contribution can and have been proposed based either on a different assumption for the form of $v_{0,s}$ or taking droplet-droplet interactions into account.²¹ Typically, these alternate expressions lead to a slightly modified phase diagram not affecting the overall character of it.

1. Solubilization limit

The solubilization limit is derived in the same way as in the previous section. For the spheres alone one finds

$$\frac{1}{R_0} = \frac{1}{6} (2+x) + \frac{t}{48} [2 \ln(\phi) - 3]. \quad (3.6)$$

In the literature it is more common to assume that $v_{0,s}$ is constant and the above expression for the solubilization limit becomes²³

$$\frac{1}{R_0} = \frac{1}{6} (2+x) + \frac{t}{24} \ln \left(\frac{\phi v_{0,s}}{\frac{4}{3} \pi R_s^3} \right), \quad (3.7)$$

with $R_s = 3$.

2. Transition to the lamellar phase

The transition to the lamellar phase is simply determined by setting $F_s = 0$ in Eq. (3.5) giving

$$\frac{1}{R_0} = \frac{1}{12} (2+x) + \frac{t}{48} [\ln(\phi) - 1]. \quad (3.8)$$

B. Spherocylinders

In order to derive an expression for the entropy of the cylinder we need to take the length of the cylinder into con-

sideration. In the Flory–Huggins theory for polymers,²² the length is expressed as $L = nl_0$ with n the number of monomers with length l_0 . The entropic contribution to the free energy of N_c cylinders, all having length L , on a lattice with volume element $v_{0,c}$ is, in the Flory–Huggins mean-field approximation, given by

$$F_{\text{ent},c} = k_B T \left[N_c \ln \left(\frac{N_c v_{0,c}}{V} \right) - N_c n \right]. \quad (3.9)$$

When we consider the curvature energy of cylinders of finite length L , the curvature energy associated with the “end-caps” of the cylinder needs to be considered. A full treatment to determine the end-cap energy involves the minimization of the free energy with respect to the full shape.¹⁷ In the present treatment, however, we assume the shape to be that of a “sphero-cylinder” (spherical end-caps) and only minimize with respect to the two shape parameters defining the spherocylinder: the radius R_c and length L . The total free energy of spherocylinders with entropy is then

$$\begin{aligned} \frac{F_c}{\pi k} = N_c \left\{ -\frac{4L}{R_0} + \frac{L}{R_c} - \frac{16}{R_0} R_c + 4(2+x) \right. \\ \left. + t \left[\ln \left(\frac{N_c v_{0,c}}{V} \right) - \frac{L}{l_0} \right] \right\}, \end{aligned} \quad (3.10)$$

with the volume and surface area given by

$$\begin{aligned} A &= N_c 2\pi R_c (L + 2R_c), \\ V_{\text{tot}} &= N_c \frac{\pi}{3} R_c^2 (3L + 4R_c). \end{aligned} \quad (3.11)$$

In reality, the cylinders are not all of the same length L and one should consider the effect of polydispersity in the cylinder length.

1. Length polydispersity

In order to account for polydispersity in the cylinder length, we need to allow for a distribution $N_c(n)$ denoting the number of cylinders with length $L = nl_0$. The total free energy then becomes a *functional* of the distribution $N_c(n)$,

$$\begin{aligned} \frac{F_c}{\pi k} = \sum_n N_c(n) \left\{ -\frac{4nl_0}{R_0} + \frac{nl_0}{R_c} - \frac{16}{R_0} R_c + 4(2+x) \right. \\ \left. + t \left[\ln \left(\frac{N_c(n) v_{0,c}}{V} \right) - n \right] \right\}, \end{aligned} \quad (3.12)$$

with the volume and surface area given by

$$\begin{aligned} A &= \sum_n N_c(n) 2\pi R_c (nl_0 + 2R_c), \\ V_{\text{tot}} &= \sum_n N_c(n) \frac{\pi}{3} R_c^2 (3nl_0 + 4R_c). \end{aligned} \quad (3.13)$$

In the following we replace the summation by an integration over n . The minimization of the free energy in Eq. (3.12), taking the above constraints on the volume and surface area into account, is done in two steps. First, it is noted that the functional differentiation with respect to $N_c(n)$ yields an exponential distribution for $N_c(n)$,

$$N_c(n) = \frac{V}{v_{0,c}} e^{-\alpha n + \beta} \quad (3.14)$$

with α and β constants to be determined from a further minimization. Depending on the molecular model used, the volume $v_{0,c}$ might depend on n , so that the exponential distribution above may have an algebraic prefactor. In the present treatment we take $v_{0,c}$ independent of n .

Instead of α and β as parameters, it is more convenient²⁴ to express the exponential distribution in terms of the average length, L , and average number of cylinders, N_c , defined as

$$\begin{aligned} N_c &\equiv \int_0^\infty dn N_c(n), \\ L &\equiv \frac{1}{N_c} \int_0^\infty dn N_c(n) nl_0, \end{aligned} \quad (3.15)$$

so that the distribution [Eq. (3.14)] becomes

$$N_c(n) = \frac{N_c l_0}{L} e^{-(nl_0/L)}. \quad (3.16)$$

As a second step, we insert the above exponential distribution back into the expression for the free energy in Eq. (3.12) and carry out the integration over n .²⁴ We find an expression for the free energy quite similar to the expression for the free energy *without* length polydispersity [Eq. (3.10)],

$$\begin{aligned} \frac{F_c}{\pi k} = N_c \left\{ -\frac{4L}{R_0} + \frac{L}{R_c} - \frac{16}{R_0} R_c + 4(2+x) \right. \\ \left. + t \left[\ln \left(\frac{N_c v_{0,c} l_0}{VL} \right) - \frac{L}{l_0} - 1 \right] \right\}, \end{aligned} \quad (3.17)$$

with the boundary conditions of the same form as in Eq. (3.11),

$$\begin{aligned} A &= N_c 2\pi R_c (L + 2R_c), \\ V_{\text{tot}} &= N_c \frac{\pi}{3} R_c^2 (3L + 4R_c). \end{aligned} \quad (3.18)$$

Again, certain assumptions need to be made regarding the unknown volume element $v_{0,c}$ and the length scale l_0 , which is the length scale over which two cylinders need to differ in length in order for the two cylinders to “count” as having different lengths. Similar to the case of spherical droplets, we assume that $v_{0,c}$ is of the order of a cylindrical segment with radius R_c and length l_0 ,

$$v_{0,c} \approx \pi R_c^2 l_0, \quad (3.19)$$

and assume l_0 to be of the order of the radius of the cylinder,

$$l_0 \approx R_c. \quad (3.20)$$

Other approximations are certainly possible and one could argue that $v_{0,c}^{1/3}$ and l_0 are fixed microscopic length scales to be determined by some other method. The assumptions in Eqs. (3.19) and (3.20) have the advantage that no additional unknown parameters have to be introduced. In the context of the approximations made, this certainly suffices but below other approaches will be considered.

With the total volume fraction $\phi \equiv V_{\text{tot}}/V$, the free energy in Eq. (3.17) is written as

$$\frac{F_c}{\pi k} = N_c \left\{ -\frac{4L}{R_0} + \frac{L}{R_c} - \frac{16}{R_0} R_c + 4(2+x) + t \left[\ln \left(\frac{\phi R_c^2}{L^2} \right) - \frac{L}{R_c} - 1 \right] \right\}, \quad (3.21)$$

where we have neglected terms of $\mathcal{O}(1/L)$.

In the case that only cylinders are present, the above free energy needs to be minimized with respect to R_c , L , and N_c keeping the volume and surface area constraints in mind. One finds that R_c and L are determined by the following two equations:

$$-\frac{8R_c}{R_0} + (7+6x) - \frac{4R_c}{L} + \frac{3}{2}t \left[\ln \left(\frac{\phi R_c^2}{L^2} \right) + 1 \right] + 5t \left(1 + \frac{14}{5} \frac{R_c}{L} \right) = 0, \quad (3.22)$$

$$\frac{6}{R_c} \frac{(L+2R_c)}{(3L+4R_c)} = 1. \quad (3.23)$$

To leading order in $1/L$ these two equations are solved explicitly to yield

$$R_c = 2 + \frac{8}{3} \frac{1}{L} + \mathcal{O}\left(\frac{1}{L^2}\right), \quad (3.24)$$

$$t \ln(L) = -\frac{16}{3} \frac{1}{R_0} + \frac{1}{3}(7+6x) + \frac{t}{2} \left[\ln(4\phi) + \frac{13}{3} \right] + \mathcal{O}\left(\frac{1}{L}\right).$$

Furthermore, the solubilization limit and transition to the lamellar phase can be determined.

2. Solubilization limit

The solubilization limit is given by

$$\frac{1}{R_0} = \frac{1}{4}(1-t) + \frac{1}{L} \left(\frac{2}{3} - \frac{23}{12}t \right) + \mathcal{O}\left(\frac{1}{L^2}\right), \quad (3.25)$$

with the cylinder length L at the solubilization limit to leading order given by

$$t \ln(L) = (1+2x) + \frac{t}{2} [\ln(4\phi) + 7] + \mathcal{O}\left(\frac{1}{L}\right), \quad (3.26)$$

which is derived by inserting the expression for $1/R_0$ in Eq. (3.25) into Eq. (3.24).

3. Transition to the lamellar phase

The transition to the lamellar phase is determined by setting $F_c = 0$ in Eq. (3.21)

$$\frac{1}{R_0} = \frac{1}{8}(1-t) - \frac{t}{2L} + \mathcal{O}\left(\frac{1}{L^2}\right), \quad (3.27)$$

with to leading order in L ,

$$t \ln(L) = \frac{1}{3}(5+6x) + \frac{t}{2} \left[\ln(4\phi) + \frac{17}{3} \right] + \mathcal{O}\left(\frac{1}{L}\right), \quad (3.28)$$

which is derived by inserting the expression for $1/R_0$ in Eq. (3.27) into Eq. (3.24).

In the calculation of the two conditions in Eqs. (3.22) and (3.23), we have made use of the expressions for $v_{0,c}$ and l_0 in Eqs. (3.19) and (3.20). As an aside we investigate the consequences of assuming that $v_{0,c}$ and l_0 are constants instead. One finds that the condition in Eq. (3.22) now becomes

$$-\frac{8R_c}{R_0} + (7+6x) - \frac{4R_c}{L} + \frac{3}{2}t \left[\ln \left(\frac{\phi v_{0,c} l_0}{\pi R_c^2 L^2} \right) + 1 \right] + 4t \frac{R_c}{l_0} \left(1 + \frac{R_c}{L} + \frac{l_0}{L} \right) = 0. \quad (3.29)$$

The difference in approach only shows up as an end-correction to the last term in Eq. (3.22). As argued before, such detail is lost in the approximative scheme considered here.

Having derived the free energies of spheres and cylinders separately, it is now easy to construct the free energy of the system containing both spheres and cylinders.

C. Spheres and spherocylinders

The total free energy of spheres and spherocylinders with length polydispersity is the sum of the free energies in Eqs. (3.5) and (3.21) taking the respective volume fractions into account,

$$\begin{aligned} \frac{\omega^3 F}{k V_{\text{tot}}} = & v_s \left\{ -\frac{12}{R_0} \frac{1}{R_s^2} + 3(2+x) \frac{1}{R_s^3} + t \frac{3}{4R_s^3} [\ln(\phi v_s) - 1] \right\} \\ & + \frac{3 v_c L}{(3L+4R_c)} \left\{ -\frac{4}{R_0} \frac{1}{R_c^2} + \frac{1}{R_c^3} - \frac{16}{R_0} \frac{1}{R_c L} \right. \\ & \left. + 4(2+x) \frac{1}{R_c^2 L} + t \frac{1}{R_c^2 L} \left[\ln \left(\frac{\phi v_c R_c^2}{L^2} \right) - \frac{L}{R_c} - 1 \right] \right\}, \end{aligned} \quad (3.30)$$

with the volume and area constraints given by

$$v_s + v_c = 1, \quad (3.31)$$

$$\frac{3v_s}{R_s} + 6 \frac{v_c}{R_c} \frac{(L+2R_c)}{(3L+4R_c)} = 1.$$

The free energy above needs to be minimized with respect to the five variables R_s , R_c , L , v_c , and v_s with the two constraints in Eq. (3.31). This has been done, numerically, with the result shown in Fig. 2. In this example we have fixed the

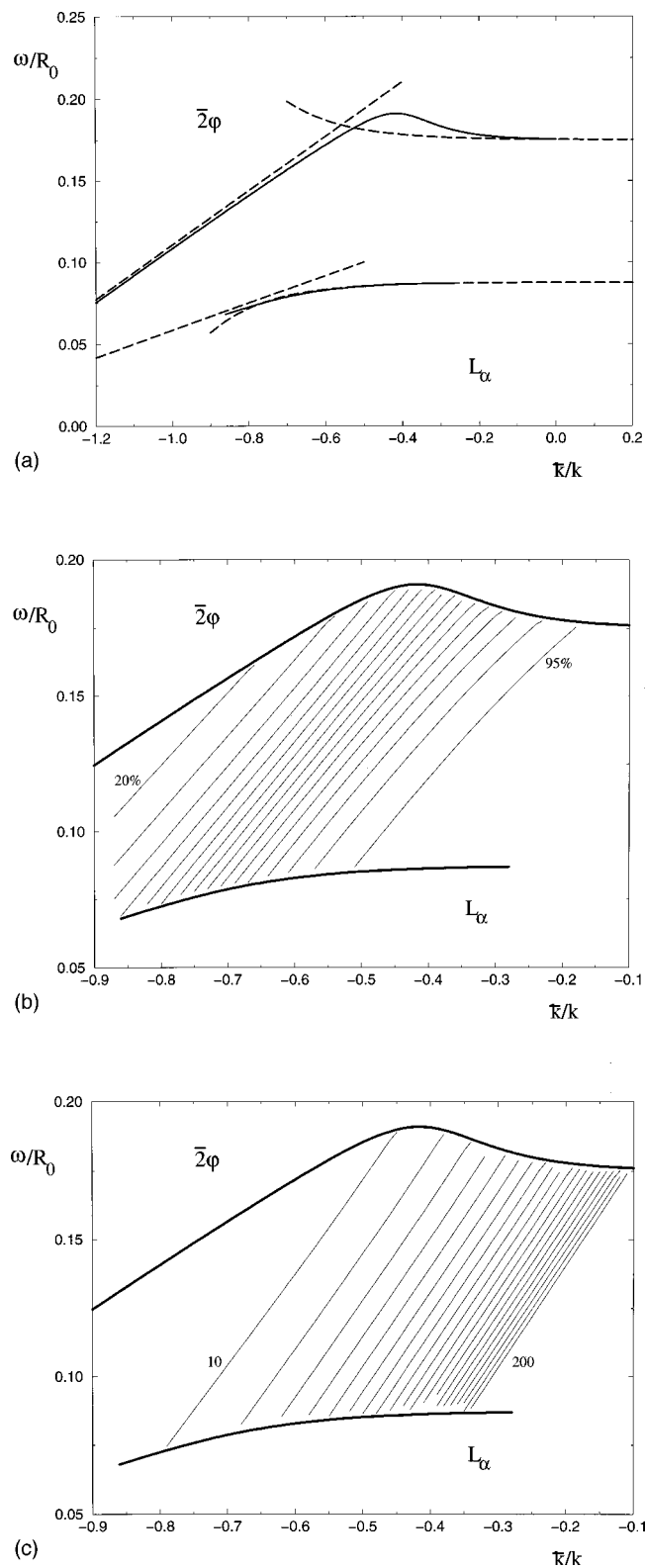


FIG. 2. Microemulsion phase diagram with translational entropy and cylinder length polydispersity, as a function of ω/R_0 and \bar{k}/k with $\phi=0.05$ and $t=0.3$. The drawn lines denote the location of the solubilization limit and the transition to the lamellar phase. In (a) the dashed lines are the limiting analytical results for the solubilization limit and transition to the lamellar phases as given in Eqs. (3.6), (3.8), (3.25), and (3.27). The cylinder volume fraction, $v_c=20\%-95\%$ (steps of 5%), and average cylinder length, $L=10-200$ (steps of 10, in units of ω) are shown in (b) and (c), respectively.

droplet volume fraction $\phi=0.05$ and set the reduced temperature $t=0.3$ (which corresponds to $k \approx 1 k_B T$). The dashed lines in Fig. 2(a) are the limiting analytical results for the solubilization limit and transition to the lamellar phases as given in Eqs. (3.6), (3.8), (3.25), and (3.27). An important distinction with the phase diagram in Fig. 1 is the fact that there is no sharp transition to a region with only spheres or only cylinders. The inclusion of translational entropy therefore smoothens the sphere to cylinder transition. This means that at any finite temperature the relative population of spheres and cylinders is determined by the Boltzmann distribution prohibiting the existence of regions with only spheres or cylinders present.

In Figs. 2(b) and 2(c), the cylinder volume fraction and cylinder length, respectively, are shown in the phase diagram. In the direction of increasing $x \equiv \bar{k}/k$ and decreasing ω/R_0 both the cylinder volume fraction and cylinder length increases.

Already the (numerical) minimization of the free energy in Eq. (3.30) gives a good indication of the influence of entropy on the phase diagram of spherical and cylindrical microemulsions. For the comparison with the experimental phase diagram, however, we still need to consider one additional effect: the polydispersity in the radius of the spheres and cylinders.

IV. RADIAL POLYDISPERSITY

In this section we account for the polydispersity in the radius of the spherical and cylindrical structures. As in the case of length polydispersity, we now have a distribution $N_s(n)$ ($N_c(n)$) denoting the number of spheres (cylinders) with radius $R_s = nr_{0,s}$ ($R_c = nr_{0,c}$).

A. Spheres

We first consider the free energy of spherical droplets,

$$\frac{F_s}{\pi k} = \sum_n N_s(n) \left\{ -\frac{16}{R_0} nr_{0,s} + 4(2+x) + t \left[\ln \left(\frac{N_s(n) v_{0,s}}{V} \right) - 1 \right] \right\}, \quad (4.1)$$

with the volume and surface area now given by

$$A = \sum_n N_s(n) 4\pi (nr_{0,s})^2, \quad (4.2)$$

$$V_{\text{tot}} = \sum_n N_s(n) \frac{4\pi}{3} (nr_{0,s})^3.$$

Again, the summation is replaced by an integration over n . The above free energy is minimized adding Lagrange multipliers σ and $-\Delta p$ fixing the surface area and volume, respectively. The distribution then has the form,

$$N_s(n) = \frac{V}{v_{0,s}} e^{-(1/t)[4(2+x) - (16/R_0)nr_{0,s} + (\sigma/k)4(nr_{0,s})^2 - (\Delta p/k)(4/3)(nr_{0,s})^3]} \quad (4.3)$$

Again, the above distribution may have an algebraic prefactor due to some assumed n (radial) dependence of $v_{0,s}$. The above distribution was first derived by Overbeek.²⁵ Following Reiss,²⁶ $v_{0,s}$ was taken proportional to $n^{-3/2}$ by Overbeek but different models leading to different exponents of the prefactor have been reported in the literature.^{27,28} Here we proceed by assuming that $r_{0,s}$ is some microscopic length scale to be determined in some other way while the expression for $v_{0,s}$ is the same as in the previous section [cf. Eq. (3.4)],

$$v_{0,s} \approx \frac{4\pi}{3} R_s^3. \quad (4.4)$$

We should now proceed in a similar way as in the treatment of the length polydispersity of the cylinder: (1) assume some form for $v_{0,s}$ such as in Eq. (4.4), (2) express the Lagrange multipliers in terms of an average radius and total number of droplets, (3) insert the resulting distribution into the free energy, and (4) minimize with respect to the remaining variables. This route is, however, mathematically rather complicated if no further approximations are made.²⁹ What we will do here is to approximate the distribution in Eq. (4.3) by discarding the terms proportional to n^2 and n^3 in the exponent and allow for the presence of some algebraic prefactor with an exponent which we will call z_s . With this approximation, the important characteristics of the distribution remain with the neglect of the n^2 and n^3 -terms in the exponent only affecting the tail-end of the distribution. The advantage of this approach is that the resulting distribution has the form of the well-known *Schultz distribution*,³⁰ widely used in the experimental fit^{7,11} of the size-distribution of microemulsion droplets,

$$N_s(n) = \frac{N_s r_{0,s}}{R_s} \frac{(z_s + 1)^{z_s + 1}}{\Gamma(z_s + 1)} \left(\frac{nr_{0,s}}{R_s} \right)^{z_s} e^{-(z_s + 1)nr_{0,s}/R_s}, \quad (4.5)$$

where $\Gamma(x)$ is Euler's Gamma function, and where N_s and R_s are the total number of droplets and average radius defined by

$$N_s \equiv \int_0^\infty dn N_s(n),$$

$$R_s \equiv \frac{1}{N_s} \int_0^\infty dn N_s(n) nr_{0,s}. \quad (4.6)$$

The constant z_s is related to the radial polydispersity, $\sigma_s^2 = 1/(z_s + 1)$. One can imagine two approaches with respect to the determination of the value of z_s ; first, z_s can be treated as a constant to be fitted to the experimental value, second, one could determine z_s from a minimization of the free energy with respect to z_s , so that z_s is expressed in terms of temperature, the total volume and surface area, and the curvature coefficients. The latter approach is more fundamental

but mathematically more complex. However, if we assume $\sigma_s^2 \ll 1$ (and later $\sigma_c^2 \ll 1$), which is usually a very good approximation, and only keep track of the leading contributions to the free energy, the minimization can be carried out analytically giving explicit expressions for the radial polydispersities.

With the distribution in Eq. (4.5) in terms of z_s and the variables N_s and R_s , insertion of $N_s(n)$ into the free energy in Eq. (4.1) and integration over n leaves us with the following expression for the free energy of polydisperse spheres:

$$\frac{F_s}{\pi k} = N_s \left\{ -\frac{16}{R_0} R_s + 4(2+x) + t \left[\ln \left(\frac{\phi r_{0,s}}{R_s} \right) + \alpha_1 \right] \right\}, \quad (4.7)$$

with the volume and area constraints given by

$$A = N_s 4\pi \alpha_2 R_s^2, \quad (4.8)$$

$$V_{\text{tot}} = N_s \frac{4\pi}{3} \alpha_3 R_s^3.$$

The functions α_1 , α_2 , and α_3 appearing in Eqs. (4.7) and (4.8) are defined as

$$\alpha_1 \equiv \ln \left(\frac{z_s + 1}{\Gamma(z_s + 1)} \right) + z_s \psi(z_s + 1) - z_s - 2$$

$$\approx -\frac{1}{2} \ln(\sigma_s^2) - \frac{3}{2} - \frac{1}{2} \ln(2\pi) + \frac{1}{3} \sigma_s^2, \quad (4.9)$$

$$\alpha_2 \equiv \frac{z_s + 2}{z_s + 1} \approx 1 + \sigma_s^2,$$

$$\alpha_3 \equiv \frac{(z_s + 2)(z_s + 3)}{(z_s + 1)^2} \approx 1 + 3\sigma_s^2,$$

with $\psi(x)$ Euler's psi function. Apart from the presence of the functions α_1 , α_2 , and α_3 , the expression for the free energy in Eq. (4.7) is the same as the free energy of the monodisperse droplets [Eq. (3.5)].

As in the previous sections, the radius of the spheres is directly determined by the constraints in Eq. (4.8) giving $R_s = 3\alpha_2/\alpha_3 \approx 3 - 6\sigma_s^2$ in units of ω . Insertion into the free energy in Eq. (4.7) and differentiation with respect to σ_s allows the determination of the radial polydispersity. One finds that in an expansion in $\sigma_s^2 \ll 1$, σ_s is determined by the following equation, which can readily be solved numerically:

$$0 = -\frac{t}{24} + \sigma_s^2 \left[-\frac{4}{R_0} + 2 + x \right] + \frac{t}{8} \sigma_s^2 \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right)$$

$$- \frac{11}{36} t \sigma_s^2 + \frac{t}{12} \sigma_s^4 \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) + \mathcal{O}(\sigma_s^4). \quad (4.10)$$

Furthermore, the solubilization limit and transition to the lamellar phase are calculated.

1. Solubilization limit

The solubilization limit is given by

$$\begin{aligned} \frac{1}{R_0} = & \frac{1}{6}(2+x) + \frac{t}{48} \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) - \frac{t}{16} \\ & + \sigma_s^2 \left[\frac{1}{3}(2+x) + \frac{t}{24} \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) - \frac{t}{36} \right] + \mathcal{O}(\sigma_s^4), \end{aligned} \quad (4.11)$$

with σ_s now determined by

$$\begin{aligned} 0 = & -\frac{t}{24} + \frac{1}{3} \sigma_s^2 (2+x) + \frac{t}{24} \sigma_s^2 \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) - \frac{t}{18} \sigma_s^2 \\ & + \frac{t}{12} \sigma_s^4 \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) + \mathcal{O}(\sigma_s^4). \end{aligned} \quad (4.12)$$

2. Transition to the lamellar phase

The transition to the lamellar phase is derived by setting $F_s = 0$ in Eq. (4.7),

$$\begin{aligned} \frac{1}{R_0} = & \frac{1}{12}(2+x) + \frac{t}{96} \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) - \frac{t}{32} \\ & + \sigma_s^2 \left[\frac{1}{6}(2+x) + \frac{t}{48} \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) - \frac{t}{72} \right] + \mathcal{O}(\sigma_s^4), \end{aligned} \quad (4.13)$$

with σ_s determined by

$$\begin{aligned} 0 = & -\frac{t}{24} + \frac{2}{3} \sigma_s^2 (2+x) + \frac{t}{12} \sigma_s^2 \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) - \frac{13}{72} t \sigma_s^2 \\ & + \frac{t}{12} \sigma_s^4 \ln \left(\frac{\phi^2 r_{0,s}^2}{18\pi \sigma_s^2} \right) + \mathcal{O}(\sigma_s^4). \end{aligned} \quad (4.14)$$

B. Spherocylinders

The inclusion of polydispersity in the distribution of the radius of the cylindrical microemulsion structures follows along the same lines as the spherical droplets. We first consider the free energy of a distribution of cylinders with radius $R_c = nr_{0,c}$,

$$\begin{aligned} \frac{F_c}{\pi k} = & \sum_n N_c(n) \left\{ -\frac{4L}{R_0} + \frac{L}{nr_{0,c}} - \frac{16}{R_0} nr_{0,c} + 4(2+x) \right. \\ & \left. + t \left[\ln \left(\frac{N_c(n) v_{0,c} l_0}{VL} \right) - \frac{L}{l_0} - 1 \right] \right\}, \end{aligned} \quad (4.15)$$

with the volume and surface area now given by

$$\begin{aligned} A = & \sum_n N_c(n) 2\pi [L nr_{0,c} + 2(nr_{0,c})^2], \\ V_{\text{tot}} = & \sum_n N_c(n) \frac{\pi}{3} [3L(nr_{0,c})^2 + 4(nr_{0,c})^3]. \end{aligned} \quad (4.16)$$

The summation is replaced by an integration over n and a Schultz-distribution is assumed for the radial distribution,

$$N_c(n) = \frac{N_c r_{0,c}}{R_c} \frac{(z_c + 1)^{z_c + 1}}{\Gamma(z_c + 1)} \left(\frac{nr_{0,c}}{R_c} \right)^{z_c} e^{-(z_c + 1) nr_{0,c}/R_c}, \quad (4.17)$$

where N_c and R_c are the total number of cylinders and average radius defined by

$$\begin{aligned} N_c & \equiv \int_0^\infty dn N_c(n), \\ R_c & \equiv \frac{1}{N_c} \int_0^\infty dn N_c(n) nr_{0,c}. \end{aligned} \quad (4.18)$$

Again an assumption needs to be made concerning the lengthscales $r_{0,c}$, $v_{0,c}$, and l_0 . We proceed by assuming that $r_{0,c}$ is a microscopic constant and $v_{0,c}$ and l_0 to be given by the previous expressions [cf. Eqs. (3.19) and (3.20)],

$$\begin{aligned} v_{0,c} & \approx \pi R_c^2 l_0, \\ l_0 & \approx R_c. \end{aligned} \quad (4.19)$$

The resulting free energy is then obtained by inserting the distribution in Eq. (4.17) into Eq. (4.15) and carry out the integration,

$$\begin{aligned} \frac{F_c}{\pi k} = & N_c \left\{ -\frac{4L}{R_0} + \alpha_4 \frac{L}{R_c} - \frac{16}{R_0} R_c + 4(2+x) \right. \\ & \left. + t \left[\ln \left(\frac{\phi R_c r_{0,c}}{L^2} \right) - \frac{L}{l_0} + \alpha_5 \right] \right\}, \end{aligned} \quad (4.20)$$

with the volume and surface area given by

$$\begin{aligned} A & = N_c 2\pi R_c (L + 2\alpha_6 R_c), \\ V_{\text{tot}} & = N_c \frac{\pi}{3} R_c^2 (3\alpha_6 L + 4\alpha_7 R_c). \end{aligned} \quad (4.21)$$

Furthermore we have defined

$$\begin{aligned} \alpha_4 & \equiv \frac{z_c + 1}{z_c} = 1 + \sigma_c^2, \\ \alpha_5 & \equiv \ln \left(\frac{z_c + 1}{\Gamma(z_c + 1)} \right) + z_c \psi(z_c + 1) - z_c - 2 \\ & \approx -\frac{1}{2} \ln(\sigma_c^2) - \frac{3}{2} - \frac{1}{2} \ln(2\pi) + \frac{1}{3} \sigma_c^2, \\ \alpha_6 & \equiv \frac{z_c + 2}{z_c + 1} \approx 1 + \sigma_c^2, \\ \alpha_7 & \equiv \frac{(z_c + 2)(z_c + 3)}{(z_c + 1)^2} \approx 1 + 3\sigma_c^2. \end{aligned} \quad (4.22)$$

In the case that only cylinders are present, the free energy in Eq. (4.20) needs to be minimized with respect to R_c , L , σ_c , and N_c keeping the volume and surface area constraints in Eq. (4.21) in mind. One finds that R_c , σ_c , and L are to leading order in $1/L$ given by

$$\begin{aligned}
R_c &= 2(1 - 2\sigma_c^2) + \frac{8}{3} \frac{1}{L} + \mathcal{O}\left(\frac{1}{L^2}\right), \\
\sigma_c^2 &= \frac{1}{L} \frac{t}{3 - \frac{8}{R_0} - 2t} + \mathcal{O}\left(\frac{1}{L^2}\right), \\
t \ln(L) &= \frac{4}{9}(7 + 6x) - \frac{64}{9} \frac{1}{R_0} \\
&\quad + \frac{t}{3} \ln\left(\frac{2\phi^2 r_{0,c}^2}{\pi t} \left(3 - \frac{8}{R_0} - 2t\right)\right) + \frac{23}{9} t + \mathcal{O}\left(\frac{1}{L}\right).
\end{aligned} \tag{4.23}$$

Furthermore, the solubilization limit and transition to the lamellar phase can be determined.

1. Solubilization limit

For a system of only cylinders one finds for the solubilization limit,

$$\frac{1}{R_0} = \frac{1}{4}(1 - t) + \frac{1}{12}(8 - 14t - 3t^2) \frac{1}{L} + \mathcal{O}\left(\frac{1}{L^2}\right), \tag{4.24}$$

with σ_c^2 and the cylinder length L to leading order given by

$$\begin{aligned}
\sigma_c^2 &= \frac{t}{L} + \mathcal{O}\left(\frac{1}{L^2}\right), \\
t \ln(L) &= \frac{4}{3}(2 + x) + \frac{t}{3} \ln\left(\frac{2\phi^2 r_{0,c}^2}{\pi t}\right) + \frac{13}{3} t + \mathcal{O}\left(\frac{1}{L}\right).
\end{aligned} \tag{4.25}$$

2. Transition to the lamellar phase

The transition to the lamellar phase for the cylinders alone is

$$\frac{1}{R_0} = \frac{1}{8}(1 - t) - \frac{3}{8} t \frac{1}{L} + \mathcal{O}\left(\frac{1}{L^2}\right), \tag{4.26}$$

with σ_c^2 and the cylinder length L to leading order given by

$$\begin{aligned}
\sigma_c^2 &= \frac{t}{2 - t} \frac{1}{L} + \mathcal{O}\left(\frac{1}{L^2}\right), \\
t \ln(L) &= \frac{4}{9}(1 + 6x) + \frac{t}{3} \ln\left(\frac{2\phi^2 r_{0,c}^2}{\pi t} (2 - t)\right) + \frac{31}{9} t \\
&\quad + \mathcal{O}\left(\frac{1}{L}\right).
\end{aligned} \tag{4.27}$$

C. Spheres and spherocylinders

Finally, we introduce the volume fractions v_s and v_c and obtain the free energy of the system containing both spheres and cylinders using Eqs. (4.7) and (4.20),

$$\begin{aligned}
\frac{\omega^3 F}{k V_{\text{tot}}} &= \frac{v_s}{\alpha_3} \left\{ -\frac{12}{R_0} \frac{1}{R_s^2} + 3(2 + x) \frac{1}{R_s^3} + t \frac{3}{4 R_s^3} \right. \\
&\quad \times \left[\ln\left(\frac{\phi v_s r_{0,s}}{R_s}\right) + \alpha_1 \right] \left. + \frac{3 v_c L}{(3 \alpha_6 L + 4 \alpha_7 R_c)} \right. \\
&\quad \times \left\{ -\frac{4}{R_0} \frac{1}{R_c^2} + \frac{\alpha_4}{R_c^3} - \frac{16}{R_0} \frac{1}{R_c L} + 4(2 + x) \frac{1}{R_c^2 L} \right. \\
&\quad \left. \left. + t \frac{1}{R_c^2 L} \left[\ln\left(\frac{\phi v_c R_c r_{0,c}}{L^2}\right) - \frac{L}{l_0} + \alpha_5 \right] \right\} \right\}, \tag{4.28}
\end{aligned}$$

with the volume and area constraints

$$\begin{aligned}
v_s + v_c &= 1, \\
\frac{3 \alpha_2}{\alpha_3} \frac{v_s}{R_s} + 6 \frac{v_c}{R_c} \frac{(L + 2 \alpha_2 R_c)}{(3 \alpha_2 L + 4 \alpha_3 R_c)} &= 1.
\end{aligned} \tag{4.29}$$

Besides temperature t and the curvature coefficients x and $1/R_0$, the inclusion of radial polydispersity has left us with the additional parameters $r_{0,s}$ and $r_{0,c}$. Furthermore, the volume and area constraints manifest themselves in the presence of the total volume fraction ϕ and length scale ω .

The free energy in Eq. (4.28) is expressed in terms of the seven variables R_s , R_c , L , z_s , z_c , v_s , and v_c to be determined by minimization of the free energy with the constraints in Eq. (4.29). This has been done, numerically, with the result shown in Fig. 3. In this example we have set $r_{0,s} = 1$ and $r_{0,c} = 1$ (in units of ω). Similar results are obtained when different values for $r_{0,s}$ and $r_{0,c}$ are assumed. In general, lowering r_0 shifts the phase boundaries uniformly to the right in the phase diagrams depicted in Fig. 3 ($\Delta x \approx 0.17$ per factor 10 in r_0).

The general shape of the phase diagram and the evolution of the cylinder volume fraction [Fig. 3(a)] and cylinder length [Fig. 3(b)] is the same as in Fig. 2. The advantage of including radial polydispersity therefore mainly lies in the fact that explicit values for the spherical [Fig. 3(c)] and cylindrical [Fig. 3(d)] radial polydispersities can be provided. It is concluded that σ_s and σ_c decrease going in the direction of the lamellar phase. Furthermore, the radial polydispersities also decrease in the direction of increasing $x \equiv \bar{k}/k$ with the important distinction, however, that σ_s takes on a minimum polydispersity of about 14% (for the few spherical droplets that remain in this region) while σ_c vanishes [$\propto 1/L$; see Eq. (4.23)]. An interesting experimental consequence of the results in Figs. 3(c) and 3(d) is that in the case that the inversion temperature is approached from below ($T < \bar{T}$), the radial polydispersity decreases with increasing temperature. This decrease in polydispersity with increasing temperature is then purely a result from the intricate interplay between the constraints, entropy, and curvature energy.

We now show how the results of this section can be compared to experimental phase diagrams. To make this comparison more transparent it should be reminded that we can take k and \bar{k} as approximately constant (fixed x and t) and treat $1/R_0$ as the “temperature variable” $1/R_0 \propto (T - \bar{T})$. In Fig. 4(a) the experimental microemulsion phase diagram is shown as a function of the oil to water ratio $\alpha \equiv o/w$

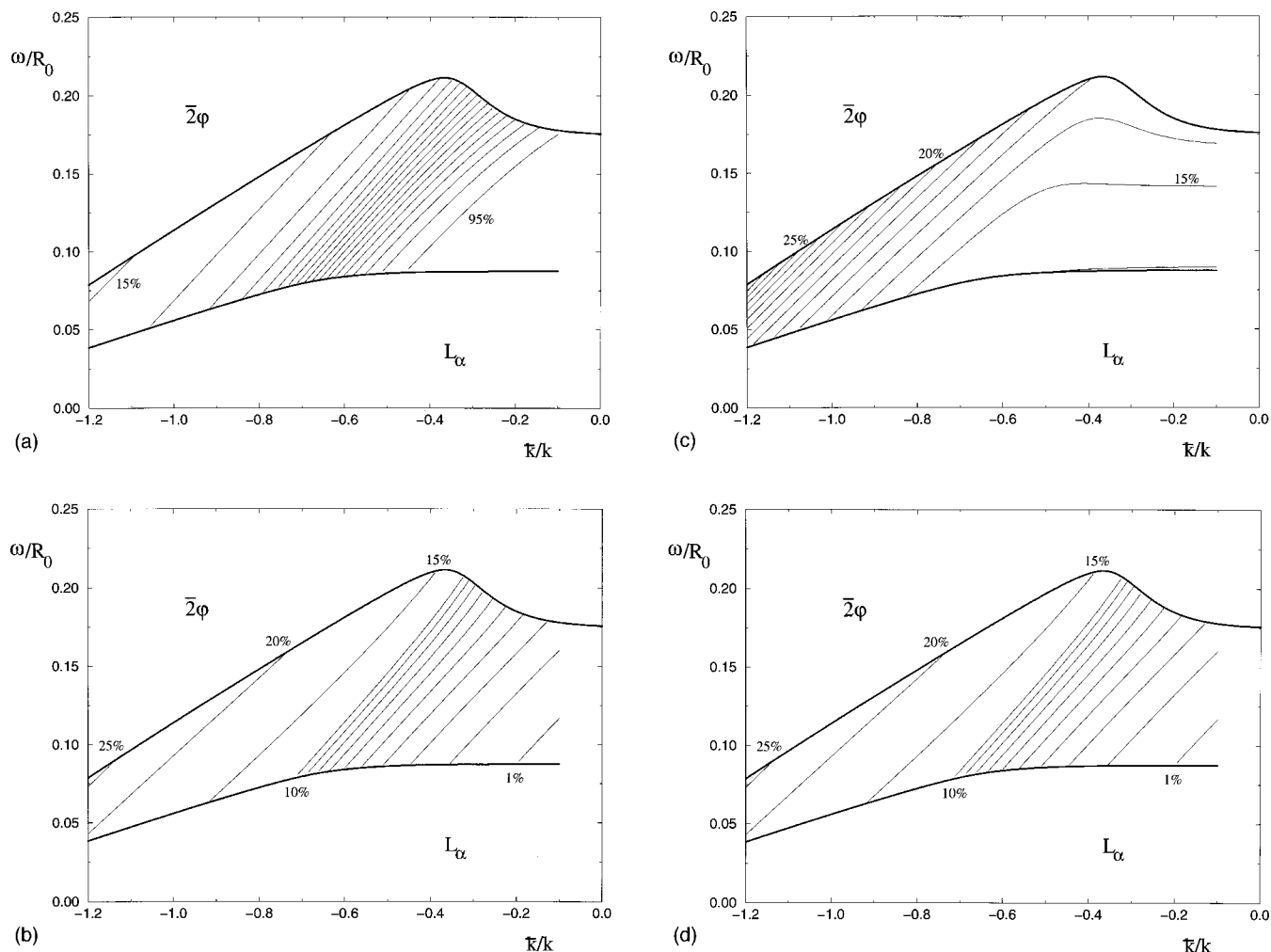


FIG. 3. Microemulsion phase diagram with translational entropy, cylinder length polydispersity, and radial polydispersities, as a function of ω/R_0 and k/k . We have chosen $\phi=0.05$, $t=0.3$ and $r_{0,s}=r_{0,c}=1$ (in units of ω). (a) shows the cylinder volume fraction, $v_c=15\%–95\%$ in steps of 5%; (b) shows the average cylinder length, $L=3–3000$ (in units of ω) in logarithmic steps; (c) shows the radial polydispersity of the sphere, $\sigma_s=14\%–27\%$ in steps of 1%; (d) shows the radial polydispersity of the cylinder, $\sigma_c=1\%–10\%$ in steps of 1% and $\sigma_c=10\%–25\%$ in steps of 5%.

+o) and temperature T at constant surfactant concentration $\gamma \equiv s/(w+s+o)$, with w , s , and o the water, surfactant, and oil weight fraction, respectively. The usual structural evolution within the one-phase region, bounded by lamellar phases (L_α) and two two-phase regions ($\bar{2}$ and $\underline{2}$), is sketched. The evolution shown is from spherical water droplets in oil (lower right corner) via a bicontinuous phase to oil droplets in water (upper right corner).³¹ The region where the results of this section are expected to be most applicable is the region close to the droplet region, not too close to the lamellar region and not too close to the bicontinuous region. We have calculated the evolution of the cylinder volume fraction taking $t=0.3$, $r_{0,s}=r_{0,c}=1$ (in units of ω) and $\gamma=0.2$ in Fig. 4(b) which roughly corresponds to the region enclosed by the dashed line in Fig. 4(a). In comparison with the usual sketch of the structural evolution, which has emerged on the basis of extensive experimental effort,³¹ it is noted that an increase in the number of cylinders (as well as average length) more prominently occurs in the direction of increasing temperature ($1/R_0 \rightarrow 0$) than with decreasing α .

In Fig. 5(a), the microemulsion phase diagram is shown

as a function of surfactant concentration γ and temperature T at constant oil to water ratio $\alpha=0.5$. This is the Kahlweit “fish”-diagram² showing a three-phase region (3) and two two-phase regions ($\bar{2}$ and $\underline{2}$). Again, the evolution is sketched from spherical water droplets in oil via a bicontinuous phase to oil droplets in water.³² The evolution of the cylinder volume fraction, taking $t=0.3$ and $r_{0,s}=r_{0,c}=1$ (in units of ω), is shown in Fig. 5(b) which roughly corresponds to the region enclosed by the dashed line in Fig. 5(a). It should be noted, however, that $\alpha=0.5$ corresponds to a rather substantial droplet volume fraction ϕ violating the assumption $\phi \ll 1$, so that the comparison between Figs. 5(a) and 5(b) should be taken only as a qualitative comparison.

V. SUMMARY AND DISCUSSION

We have showed that the Helfrich free energy model can be used to describe the sphere to cylinder transition in a one-phase region microemulsion system. In order for our description to be as realistic as possible we have included, be-

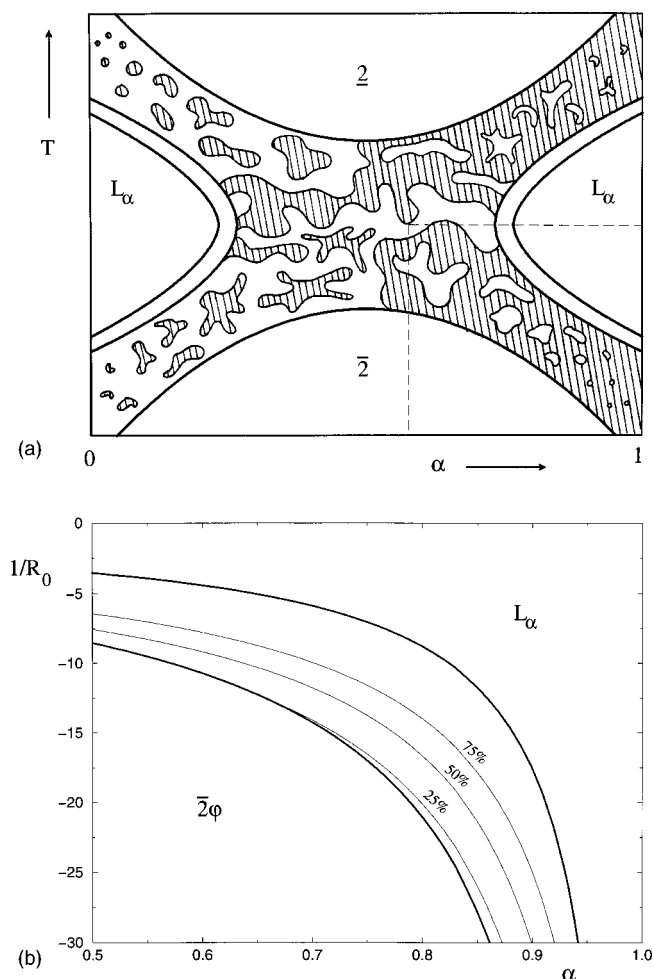


FIG. 4. Microemulsion phase diagram as a function of the oil to water ratio α and (a) temperature or (b) $1/R_0$, at fixed surfactant concentration γ . In (a) the usual structural evolution within the one-phase region is sketched. The one-phase region is bounded by two lamellar phase regions (L_α) and two two-phase regions ($\bar{2}$ and $\bar{2}$). The region enclosed by the dashed lines shows the region where the theory is expected to be most applicable. This region is shown in (b) which was calculated taking $t=0.3$, $r_{0,s}=r_{0,c}=1$ (in units of ω) and $\gamma=0.2$. Also shown is the cylinder volume fraction, $v_c=25\%–75\%$ in steps of 25%.

sides translation entropy, cylinder length polydispersity, and radial polydispersity. The model presented here has a number of limitations that should be discussed.

(i) We have neglected all interactions between the different structures present. This effectively means that the validity of our analysis is limited to small volume fractions, $\phi \ll 1$. For a realistic description at higher volume fractions, the sphere–sphere, sphere–cylinder, and cylinder–cylinder interaction energy needs to be included. Treating the spherical droplets as hard spheres or sticky hard spheres,⁸ which may or may not be a valid assumption,³³ the sphere–sphere interaction energy can be well described by a Percus–Yevick approximation, but little is known about the other interactions thus seriously hampering the extension to higher volume fractions.

(ii) Another drawback of the model is the neglect of the possible existence of other phases. In the direction of zero spontaneous curvature ($1/R_0 \rightarrow 0$) it is expected that the cylinders formed start to branch¹⁷ when they increase in length

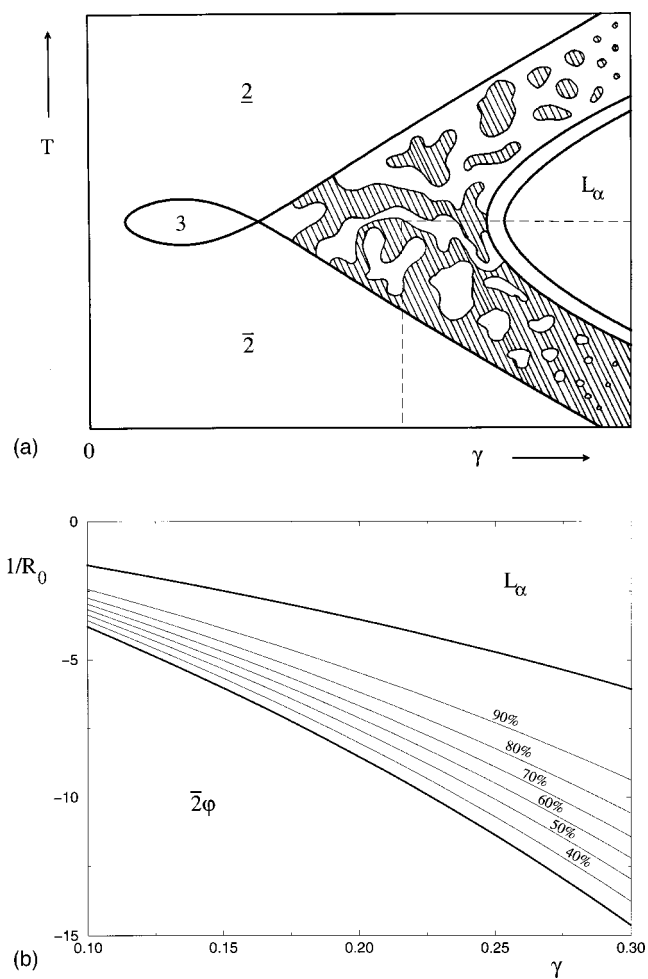


FIG. 5. Microemulsion phase diagram as a function of surfactant concentration γ and (a) temperature or (b) $1/R_0$, at fixed oil to water ratio, $\alpha=0.5$. In (a) the usual structural evolution within the one-phase region is sketched. This is the “fish”-diagram showing a three-phase region (3) and two two-phase regions ($\bar{2}$ and $\bar{2}$). The region enclosed by the dashed lines shows the region where the theory is expected to be most applicable. This region is shown in (b) which was calculated taking $t=0.3$ and $r_{0,s}=r_{0,c}=1$ (in units of ω). Also shown is the cylinder volume fraction, $v_c=40\%–90\%$ in steps of 10%.

and ultimately form a bicontinuous phase that competes with the formation of a lamellar structure. At present we have limited ourselves to the calculation of the point where the free energy changes sign, and the lamellar phase is formed, hereby neglecting the narrow region of microemulsion–lamellar phase coexistence or entropy considerations for the lamellar phase. It should therefore be concluded that our theoretical analysis is most valid close to the solubilization limit describing the onset of the sphere to cylinder transition.

(iii) The theory presented here has a mean-field character. This means that even though certain fluctuations around the mean are taken into consideration—for example those fluctuations that only change the radius—*shape fluctuations* are not taken into account. The result is that our calculated values for the radial polydispersity are a lower limit to the experimental value. Especially in the case of the radial polydispersity of very long cylinders, the contribution to the radial polydispersity of a uniform fluctuation changing only the radius of the cylinder becomes negligible compared to

the contribution from undulatory shape fluctuations.

(iv) Many assumptions and approximations have been made with regard to the consideration of entropy-effects. These include the Flory–Huggins approximation for the cylinder entropy and the assumptions made on the magnitude of the “entropy length scales” $v_{0,s}$, $v_{0,c}$, l_0 , $r_{0,s}$, and $r_{0,c}$. Although we have not shown it in great detail, it turns out that details concerning the choice of the entropy terms only deform the phase diagram slightly without changing the overall character of it. The same holds for the influence on the calculation of the structural parameters, v_c , R_s , R_c , σ_s , and σ_c . Only the average cylinder length L turns out to be logarithmically sensitive to the details of the model and the numbers presented here should therefore only be taken as an indication.

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