

Soret effect of mono- di- and tri-glycols in ethanol

Michael Klein* and Simone Wiegand

Received Xth XXXXXXXXXXXX 20XX, Accepted Xth XXXXXXXXXXXX 20XX

First published on the web Xth XXXXXXXXXXXX 200X

DOI: 10.1039/b000000x

We employed thermal diffusion forced Rayleigh scattering (TDFRS) to investigate the chain length dependence of the thermal diffusion behavior of short glycols in ethanol. We studied three glycols, monoethylenglycol (MEG), diethylenglycol (DEG) and triethylenglycol (TEG) in an ethanol solution. In contrast to the polymer polyethylenglycol, the shorter glycols used within the framework of this study are completely soluble in ethanol, at least for low molar fractions. In order to guarantee a good solubility of the glycols, the measurements were limited to molar fractions of 10 % and 20 %. As the different glycols only differ by their chain length it will be possible to deduce its influence on the thermal diffusion properties. Due to the fairly short persistence length of the glycols, D_T reaches already for the dimer a plateau value. This finding agrees with simulation results. Furthermore, the thermal expansion and the kinematic viscosity are measured to investigate if empirical correlations seen for other systems could be confirmed.

1 Introduction

Although there is large activity within the research community, the Ludwig-Soret effect, a non-equilibrium process taking place when initially at least binary mixtures are exposed to a temperature gradient, is still not understood. Due to a temperature gradient a mass flow, leading to a concentration gradient is initiated. For binary mixtures this can be described by the Soret-coefficient $S_T = D_T/D$, which is defined as the ratio between the thermal diffusion coefficient D_T and the mass diffusion coefficient D . From the applications side of view the effect is interesting for example in the fields of biology^{1,2}, energy³ and oil industry⁴.

For many binary liquid mixtures, many experimental, theoretical and simulation attempts to explain the mechanism of thermodiffusion have been made^{5–14}. A relation of the governing coefficients with physical properties such as density¹³, molecular weight¹⁴, moment of inertia⁵, viscosity¹⁵, thermal expansion¹⁵, molecular structure¹¹, and also chemical interactions¹² have been proposed. Apart from low molecular weight mixtures also polymers solutions and colloidal suspensions in polar and non-polar solvents have been investigated^{16–18}. For this study, the thermal expansion coefficient α and the kinematic viscosity ν are measured to compare to recent results on alkanes¹⁵ and sugars^{19,20}.

A few experimental, theoretical and simulation studies have been performed to investigate the chain length dependence of D_T . Based on an early work by Brochard and de Gennes²¹, Würger²² showed that the solvent molecular structure and the solute Brownian motion give rise to a contribution to the overall thermophoretic velocity. He states that this value is al-

ways negative. This model leads to an increase of D_T with increasing chain length and a plateau value for higher chain lengths. This approach describes n-alkanes and polystyrene as nicely shown by Stadelmaier *et al.*²³. Nevertheless a study by Blanco *et al.* on oligosaccharides²⁰ shows exactly the opposite trend: The thermal diffusion D_T decreases with increasing chain length until it saturates. The saturation behavior has also been found in simulations. Zhang and Müller-Plathe²⁴ showed by reverse nonequilibrium molecular dynamics simulations, employing a simple bead-spring model that the chain length for which the plateau value is reached, depends on the flexibility of the chains. In the simulations, more flexible chains reach a constant D_T at shorter chain lengths compared to rigid chains. Describing the chain length in terms of the persistence length l_p , D_T becomes constant when the chain length is of around two to three times l_p . The chain length dependence of D_T of alkanes¹⁴, polystyrene²³, sugars²⁰ and also of glycols in aqueous solution²⁵ is already documented in literature. If we follow the argumentation of the simulation results we expect for glycols in ethanol a different behavior compared to the aqueous solutions, because in contrast to water no hydration layer is formed. This should lead to a higher flexibility of the molecules and a shorter persistence length and therefore D_T should reach its plateau value already for shorter chains.

An additional motivation to study the chain length dependence of glycols came from a study of polyethylenglycol (PEG) in ethanol water mixtures²⁶. As illustrated in Fig. 1 the system shows an interesting behavior at a water weight fraction between $w = 5 - 70$ %. In this range PEG and ethanol enrich at the warm side, while between $w = 70 - 82$ % PEG goes to the warm side and ethanol to the cold. For higher water contents both go to the cold side²⁶. This makes it extremely

^a Forschungszentrum Juelich GmbH, IFF - Soft Matter, D-52428 Juelich

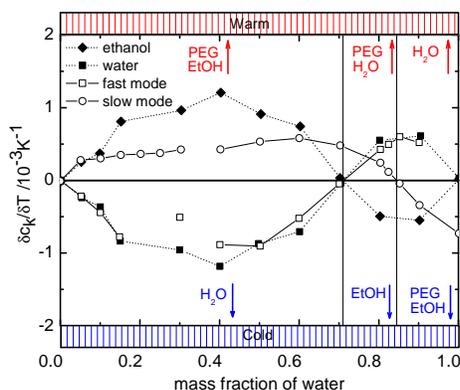


Fig. 1 Mass fraction differentials $\delta c/\delta T$ of the binary mixtures water/ ethanol, $c_w, 1 - w$, and the ternary mixtures PEO/water/ethanol, $c, c_w, 1 - c - w$, as a function of solvent composition. Closed and open symbols represent data for the binary and ternary systems, respectively. The lines are guides to the eye. Data from²⁶.

difficult to perform convection free experiments. In the case of a binary mixture it is usually possible to find a situation that the denser component moves to the bottom, but in the case of a ternary mixture with the two sign change concentration it will be impossible. This behavior makes this mixture in general extremely appealing as a benchmark system for micro-gravitational experiments. One drawback is the slow diffusing polymer, which might decompose upon longtime storage during space experiments and which also requires very long measurement times, which are often not available. Therefore it is interesting to investigate shorter glycols, which have a better longtime stability and require shorter measurement times.

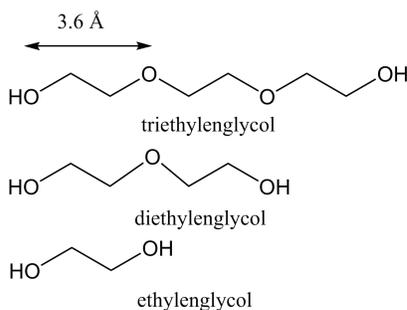


Fig. 2 Structure of the investigated glycols.

To investigate the chain length dependence of the glycols, solutions of mono-, di- and triethylenglycol (see Fig. 2 in ethanol) were measured. To improve the signal to noise ratio the regime of measurements was limited to molar concentra-

tions of 10 % and 20 % and a temperature regime from 30 °C to 50 °C.

2 Experimental section

2.1 Thermal Diffusion Forced Rayleigh Scattering

The main experimental technique employed in this study was Thermal Diffusion Forced Rayleigh Scattering (TDFRS). This technique is intensively described elsewhere^{27–29}. The setup uses a blue laser ($\lambda_w = 488\text{nm}$), which is splitted into two writing beams of equal intensity, which are aligned in such a way that they interfere within the sample cell. Thus, they form an intensity grating, which is due to a small amount of dye Quinizarin[®] in the sample solution, converted into a temperature grating. This is in turn also causing a concentration grating by the effect of thermal diffusion. Both gratings add up to a refractive index grating which is detected by a third laser beam ($\lambda_r = 633\text{nm}$). For the determination of the transport coefficients, the measured heterodyne signal is fitted using the independently measured contrast factors $(\partial n/\partial c)_{p,T}$ and $(\partial n/\partial T)_{p,c}$.

2.2 Materials and Equipment

Ethylenglycol ($\geq 99.8\%$), diethylenglycol ($\geq 99\%$), triethylenglycol ($\geq 99\%$) were purchased at Sigma Aldrich. Ethanol of 99.9 % purity by Merck was used. The mass concentrations were prepared employing a Mettler-Toledo digital scale with an accuracy of 0.0001 g. To convert the light energy into heat energy, Quinizarin[®] was added to the solution. An optical quartz cell by Hellma with an optical path length of about 0.2 mm was filled through a filter (Spartam 13/0.45 RC). An Anton Paar automatic Microviscosimeter AMVn was used to determine the kinematic viscosity of the mixtures. To determine the density ρ and the thermal expansion coefficient α an Anton Paar DMA 4500 densimeter was used. It has an accuracy of $\Delta\rho = \pm 0.00001\text{ g cm}^{-3}$ and a temperature control of $\Delta T = \pm 0.01\text{ K}$. Employing an Anton Paar RXA 156 refractometer the refractive index as function of mass concentration has been determined. The accuracy is of $\Delta n = \pm 0.00003$ with a temperature control of $\Delta T = \pm 0.01\text{ K}$. For each solution five different mass concentrations have been prepared and measured. By linear interpolation the increment in refractive index $\partial n/\partial c$ was determined. The refractive index increment with temperature $(\partial n/\partial T)$ has been measured employing a Michelson interferometer in a temperature range 2K below and above the desired temperature.

2.3 Experiments and Discussion

The glycols studied in this work have been prepared as 10 % and 20 % molar solutions. The temperature has been varied

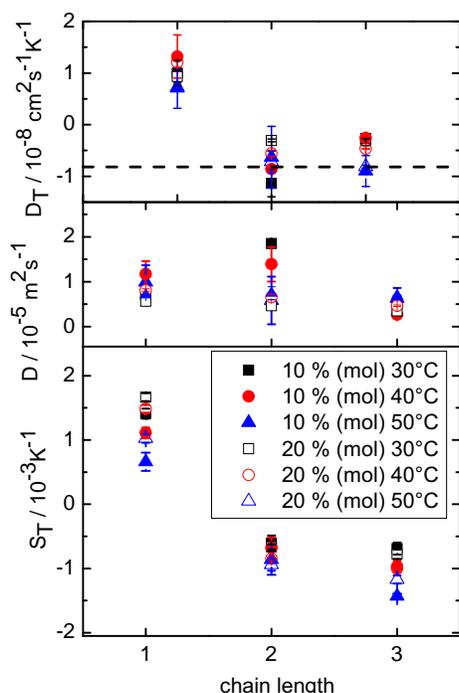


Fig. 3 Thermal diffusion coefficient D_T , mass diffusion D and Soret coefficient S_T plotted as function of chain length. It can be seen that the sign of D_T and S_T changes between the monomer and the dimer. The different molar concentrations of 10 % (closed symbols) and 20 % (open symbols) provide results which are equal within error. The dashed line in the D_T plot denotes the value of polyethyleneglycol in solution with ethanol and water as reported in²⁶.

between 30 °C and 50 °C. In Fig. 3 the chain length dependence of the thermal diffusion coefficient D_T , the mass diffusion coefficient D and the Soret coefficient S_T are plotted. It can be seen that for all variations of the two parameters concentration and temperature D_T follow roughly the same trend. Between the monomer and the dimer the sign of the coefficient changes. The values for the two different concentrations are equal within error. For the measurement series at 30 °C and 40 °C the value of D_T increases slightly between the dimer and the trimer. Already the value of D_T for the dimer is close to the value of the polymer PEO in ethanol with little addition of water²⁶, indicated by the dashed line in the plot. Simulations by Zhang *et al.*²⁴ show that the plateau in the thermal diffusion coefficient is usually reached at chain lengths of around two to three times the persistence length. The persistence length, l_p , of PEO in water can be calculated from the characteristic ratio, $C_\infty = 5.2$ determined by Kawaguchi *et al.*³⁰ using the well known expression $l_p = (b/2)(C_\infty + 1)$ given in textbooks³¹.

Assuming a bond length of $b = 1.54 \text{ \AA}$ we find $l_p = 4.8 \text{ \AA}$, which corresponds to $4/3 \approx 1.33$ monomers, if we assume that a monomer can be characterized by the oxide to oxide length of roughly 3.6 \AA . That this is a reasonable assumption has also been experimentally confirmed³². A comparison with data of glycols measured in aqueous solution by Chan *et al.* is shown in Fig. 4. It can be seen that the plateau of D_T for the aqueous solution is approached between three and four monomers, which corresponds to an approximate length of $10.8 - 11.4 \text{ \AA}$, whereas two to three times l_p is equivalent to $9.6 - 11.4 \text{ \AA}$. Thus, the trend of the simulation data is met by the experimental data. For glycols in ethanol the plateau is already reached for the dimer. This result is reasonable because in ethanol there will be no hydration layer formed, hence, possibly reducing the persistence length to one monomer. Due to the insolubility of the higher molecular weight polyethyleneglycols in ethanol no experimental data are available. The difference between the behavior in water and in ethanol might be explained by the following. In water the chain length behavior is governed by the strength of the hydrogen bonds. With increasing chain length of the glycol at constant molar fraction the hydrogen bonding gets weaker, hence leading to an increasing Soret coefficient, similar to the temperature behavior reported for PNIPAm by Kita *et al.*³³. In ethanol the behavior might be governed mainly by the quality of the solvent. The longer the chains of glycol get, the worse is ethanol as a solvent, resulting finally in the polymer which is not soluble at all in pure ethanol. Hence, the Soret coefficient decreases with increasing chain length.

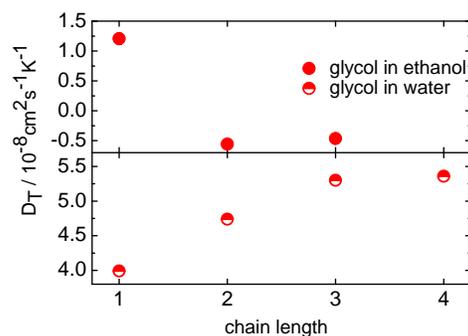


Fig. 4 Thermal diffusion coefficient D_T as function of chain length. The upper part shows D_T of glycols in ethanol measured in this work, the lower part data from Chan *et al.*³⁴ for different glycols in water. Both measurements series were carried out at 20 % molar solution at 40 °C.

The mass diffusion coefficient D , as shown in the middle of Fig. 3, shows in all cases, except the 10 % molar solution at 30 °C, a linear behavior. For the Soret coefficient S_T , seen in

the upper part Fig. 3 the situation is similar to the behavior of D_T . The errors are comparably small and the values for the two concentrations are again equal within error. A closer look at the temperature dependence yields that the value of S_T is slightly decreasing with higher temperature.

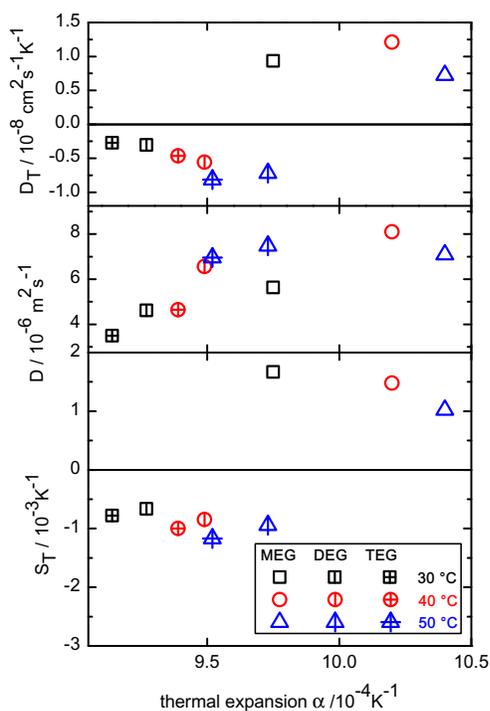


Fig. 5 Soret coefficient S_T , mass diffusion coefficient D and thermal diffusion coefficient D_T as function of the thermal expansion α for the 20 % solution. For all three temperatures S_T increases almost linearly with α , D and D_T also increase with α . The symbols with no, one and two bars correspond to mon- di and tri-mer, respectively.

In Fig. 5 the diffusion coefficient D_T , the mass diffusion coefficient D and the Soret coefficient S_T are plotted over the thermal expansion α . D and S_T clearly increase with increasing α , although the data is not as linear as for previous data on saccharides for example²⁰. It may be possible that the non-linearity can be attributed the sign change of D_T and S_T between the monomer and the dimer, which is reported for the before mentioned data.

Also the plot of the thermal diffusion coefficient D_T over the ratio of thermal expansion, α , and kinematic viscosity ν , as shown in Fig. 6 gives no better linearity. This results are in contrast to the results for alkanes¹⁵ and sugars^{19,20}. For the non-polar as well as for the polar aqueous system an excellent linear dependence of D_T with α/ν was found. We can only speculate, why this empirical correlation, which is also sup-

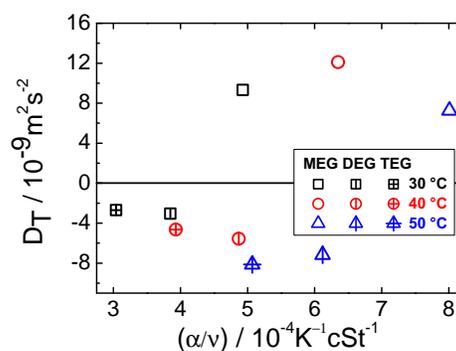


Fig. 6 Thermal diffusion D_T plotted over the ratio of thermal expansion α over the kinematic viscosity ν for the 20 % molar solution. Although the increase in thermal diffusion seems not to be linear, in general it increases with increasing ratio of α and ν . The symbols with no, one and two bars correspond to mon- di and tri-mer, respectively.

ported by theoretical arguments³⁵, fails for this aqueous system. The most striking difference compared to the previous investigated systems, is the occurrence of a sign change between the glycol and diethylen glycol. Often sign changes are related to a structural change in the solutions such as a break down of a hydrogen bond network^{26,36}. This will certainly also change the dependence on the thermal expansion coefficient. The measurements for the diethylen- and triethylen glycol indicate that for higher glycols a linear dependence might be expected, but due to the limited solubility of the higher glycols in ethanol experiments become difficult and for very higher molar masses impossible.

3 Conclusions

In this paper we presented the thermal diffusion coefficients of monoethylenglycol, diethylenglycol and triethylenglycole in pure ethanol. We found that the Soret coefficients and the thermal diffusion coefficients strongly depend on the chain length of the glycol. The D_T values coincide already between the dimer and the trimer with the value of the corresponding polymer PEG. Due to the short persistence length of the system of the order of one monomer the plateau value of D_T is already reached for the dimer, which is in accordance with the prediction of molecular dynamic simulation results²⁴. We were not able to confirm the empirical correlation between the thermal diffusion coefficient and the ratio of the thermal expansion coefficient and the kinematic viscosity. We suspect that the sign change of the Soret coefficient between the monomer and the dimer at constant molar fraction might be responsible

for this behavior. Interestingly, neither the temperature nor the concentration of the prepared solutions seems to have a large impact on the measured coefficients, at least in the regime where the measurements were conducted. This is extremely appealing for longtime experiments or experiments for which the samples have to be prepared long before the measurements, as slight or even rather big deviations of the concentration of a given solution, due to evaporation of one component yields only small changes in the measured result.

4 Acknowledgments

We would like to thank Pablo Blanco and Bastian Arlt for many fruitful discussions and Hartmut Kriegs for technical support. Financial support due to the Deutsche Forschungsgemeinschaft grant Wi 1684 is gratefully acknowledged.

References

- 1 D. Braun and A. Libchaber, *Phys. Rev. Lett.*, 2002, **89**, 2–5.
- 2 A. Bahat, I. Tur-Kaspa, A. Gakamsky, L. Giojalas, C. H. Breitbart and M. Eisenbach, *Nat. Med.*, 2003, **9**, 1–2.
- 3 C. Angeli and E. Leonardi, *Int. J. Heat Mass Transfer*, 2005, **48**, 4633–4639.
- 4 F. Montel, G. Caillet, A. Pucheu and J. Caltagirone, *Mar. Pet. Geol.*, 1993, **10**, 51–57.
- 5 C. Debuschewitz and W. Köhler, *Phys. Rev. Lett.*, 2001, **87**, 3–6.
- 6 B.-J. de Gans, R. Kita, S. Wiegand and J. Luettmmer Strathmann, *Phys. Rev. Lett.*, 2003, **91**, 245501.
- 7 S. N. Semenov and M. E. Schimpf, *Phys. Rev. E*, 2005, **72**, Art. No. 041202.
- 8 G. Galliero, M. Bugel, B. Duguay and F. Montel, *J. Non-Equilib. Thermodyn.*, 2007, **32**, 251–258.
- 9 P. A. Artola, B. Rousseau and G. Galliero, *J. Am. Chem. Soc.*, 2008, **130**, 10963–10969.
- 10 M. Hartung, J. Rauch and W. Khler, *J. Chem. Phys.*, 2006, **125**, 214904.
- 11 P. Polyakov, J. Luettmmer-Strathmann and S. Wiegand, *J. Phys. Chem. B*, 2006, **110**, 26215–26224.
- 12 G. Wittko and W. Köhler, *Europhys. Lett.*, 2007, **78**, 46007.
- 13 P. Blanco, M. Bou-Ali, J. Platten, J. Madariaga, P. Urteaga and C. Santamaria, *J. Non-Equilib. Thermodyn.*, 2007, **32**, 309–317.
- 14 P. Blanco, P. Polyakov, M. M. Bou-Ali and S. Wiegand, *J. Phys. Chem. B*, 2008, **112**, 8340–8345.
- 15 P. Blanco, M. Mounir Bou-Ali, J. K. Platten, P. Urteaga, J. a. Madariaga and C. Santamaria, *J. Chem. Phys.*, 2008, **129**, 174504.
- 16 J. Rauch and W. Köhler, *Macromolecules*, 2005, **38**, 3571–3573.
- 17 H. Ning, J. Buitenhuis, J. K. G. Dhont and S. Wiegand, *J. Chem. Phys.*, 2006, **125**, 204911.
- 18 M. Braibanti, D. Vigolo and R. Piazza, *Phys. Rev. Lett.*, 2008, **100**, year.
- 19 P. Blanco and S. Wiegand, *J. Phys. Chem. B*, 2010, **114**, 2807–2813.
- 20 P. Blanco, H. Kriegs, B. Arlt and S. Wiegand, *J. Phys. Chem. B*, 2010, **114**, 10740–10747.
- 21 F. Brochard and P. de Gennes, *C.R. Acad. Sci. Paris, Serie II*, 1981, **293**, 1025.
- 22 A. Würger, *Phys. Rev. Lett.*, 2009, **102**, 078302.
- 23 D. Stadelmaier and W. Köhler, *Macromolecules*, 2008, **41**, 6205–6209.
- 24 M. Zhang and F. Müller-Plathe, *J. Chem. Phys.*, 2006, **125**, 124903.
- 25 J. Chan, J. Popov, S. Kolisnek-Kehl and D. Leaist, *J. Solution Chem.*, 2003, **32**, 197–214.
- 26 R. Kita, S. Wiegand and J. Luettmmer Strathmann, *J. Chem. Phys.*, 2004, **121**, 3874–3885.
- 27 K. Thyagarajan and P. Lallemand, *Opt. Commun.*, 1978, **26**, 54–57.
- 28 W. Köhler and B. Müller, *J. Chem. Phys.*, 1995, **103**, 4367–4370.
- 29 W. Köhler, *J. Chem. Phys.*, 1993, **98**, 660–668.
- 30 S. Kawaguchi, G. Imai, J. Suzuki, a. Miyahara, T. Kitano and K. Ito, *Polymer*, 1997, **38**, 2885–2891.
- 31 H.-G. Elias, *Macromolecules: Volume 3: Physical Structures and Properties*, Wiley-VCH, Weinheim, 2008, vol. 3.
- 32 K. Kenworthy, K. Hristova, D. Needham and T. J. McIntosh, *Biophys. J.*, 1995, **68**, 1921–36.
- 33 R. Kita, P. Polyakov and S. Wiegand, *Macromolecules*, 2007, **40**, 1638–1642.
- 34 J. Chan, J. Popov, S. Kolisnek-Kehl and D. Leaist, *J. Solution Chem.*, 2003, **32**, 197–214.
- 35 H. Brenner, *Phys. Rev. E*, 2006, **74**, 036306.
- 36 H. Ning and S. Wiegand, *J. Chem. Phys.*, 2006, **125**, 221102.