



Investigation of the α -pinene photooxidation by OH in the atmospheric simulation chamber SAPHIR

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Abstract. The photooxidation of the most abundant monoterpene, α -pinene, by the hydroxyl radical (OH) was investigated at atmospheric concentrations in the atmospheric simulation chamber SAPHIR. Concentrations of nitric oxide (NO) were below 120 pptv. Yields of organic oxidation products are determined from measured time series giving values of 0.11 ± 0.05 , 0.19 ± 0.06 , and 0.05 ± 0.03 for formaldehyde, acetone, and pinonaldehyde, respectively. The pinonaldehyde yield is at the low side of yields measured in previous laboratory studies, ranging from 0.06 to 0.87. These studies were mostly performed at reactant concentrations much higher than observed in the atmosphere. Time series of measured radical and trace-gas concentrations are compared to results from model calculations applying the Master Chemical Mechanism (MCM) 3.3.1. The model predicts pinonaldehyde mixing ratios that are at least a factor of 4 higher than measured values. At the same time, modeled hydroxyl and hydroperoxy (HO₂) radical concentrations are approximately 25 % lower than measured values. Vereecken et al. (2007) suggested a shift of the initial organic peroxy radical (RO₂) distribution towards RO₂ species that do not yield pinonaldehyde but produce other organic products. Implementing these modifications reduces the model–measurement gap of pinonaldehyde by 20 % and also improves the agreement in

modeled and measured radical concentrations by 10 %. However, the chemical oxidation mechanism needs further adjustment to explain observed radical and pinonaldehyde concentrations. This could be achieved by adjusting the initial RO₂ distribution, but could also be done by implementing alternative reaction channels of RO₂ species that currently lead to the formation of pinonaldehyde in the model.

1 Introduction

Approximately 1000 Tg of carbon from biogenic volatile organic compounds (BVOCs) is emitted every year into the atmosphere (Guenther et al., 2012). The majority of these compounds is isoprene (53 %) followed by monoterpene species (16 %). Within the group of monoterpenes α -pinene is the most abundant species with a contribution of 6.6 % to the global emission of BVOCs. During daytime, the prevalent sinks of these compounds are ozonolysis reactions and the reaction with photochemically formed hydroxyl radicals (OH) (Calogirou et al., 1999; Atkinson and Arey, 2003) producing organic peroxy radicals (RO₂). OH is reformed in a radical reaction chain that involves reactions with nitric oxide (NO), thereby producing NO₂. This radical reaction cycle

impacts air quality because (1) the subsequent photolysis of NO_2 is the only chemical source for tropospheric ozone (O_3) and (2) oxygenated volatile organic compounds (OVOCs) are formed, which can be precursors for the formation of secondary organic aerosols (SOAs) (Glasius and Goldstein, 2016).

Field studies conducted in forested environments, which were characterized by large BVOC emissions and low NO concentrations, showed large discrepancies between measured OH radical concentrations and predictions of model calculations (e.g., Lelieveld et al., 2008; Hofzumahaus et al., 2009; Whalley et al., 2011). Under these conditions, it is expected that radical recycling is suppressed due to the dominance of radical termination reactions such as the reaction of RO_2 with hydroperoxy radicals (HO_2). Recent theoretical and laboratory studies of the chemistry of isoprene, which was often an important OH reactant in these field experiments, however, revealed that unimolecular RO_2 reactions that efficiently reform OH can compete with the reaction of RO_2 and NO for these conditions (Peeters et al., 2009, 2014; Crounse et al., 2011, 2012; Fuchs et al., 2013, 2014).

In contrast to isoprene, radical recycling in the chemistry of monoterpenes is less well investigated for conditions where field measurements indicate missing OH productions. Compared to isoprene, the degradation chemistry of monoterpene species is more complicated due to their more complex structure leading to a higher number of possible reactions and products. Laboratory and theoretical studies of the OH oxidation of α - and β -pinene focused on product yields and experiments were often performed at high reactant concentrations (see for example Vereecken et al., 2007; Vereecken and Peeters, 2012; Eddingsaas et al., 2012). The main products of the α -pinene + OH photooxidation are pinonaldehyde, acetone, and formaldehyde (HCHO). Product yields determined in previous laboratory studies were highly variable. For example, pinonaldehyde yields ranged from 6 % to 87 % (Larsen et al., 2001; Nozière et al., 1999).

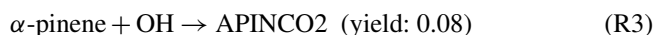
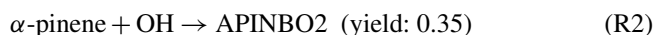
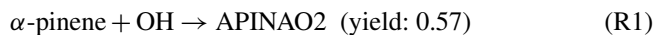
In two field studies in environments in which monoterpenes and 2-methyl-3-buten-2-ol (MBO) were the most important biogenic organic compounds (Kim et al., 2013; Hens et al., 2014) missing OH production in model calculations was found. In addition, HO_2 concentrations were underestimated. If the model was constrained to measured HO_2 , model–measurement discrepancies in OH became small due to the enhanced OH production in the reaction of HO_2 with NO. This indicated that the chemical system of monoterpenes as currently implemented in models lacks an HO_2 source. A chamber study investigating the OH oxidation of β -pinene gave similar results (Kaminski et al., 2017). Another chamber study looking at the OH oxidation of MBO showed that OH recycling in the MBO chemistry is well understood, indicating that monoterpenes are responsible for the missing HO_2 in the field campaigns (Novelli et al., 2018).

In this study, the photooxidation of α -pinene by OH was investigated in experiments in the atmospheric simulation

chamber SAPHIR (Simulation of Atmospheric Photochemistry In a large Reaction chamber) at Forschungszentrum Jülich. Experiments were performed under controlled and atmospherically relevant conditions found in forested environments with NO mixing ratios less than 120 pptv. Aerosol formation did not play a role in these experiments because no significant nucleation was observed. Measured time series were compared to model results from the Master Chemical Mechanism in the recent version 3.3.1 (MCM, 2019; Jenkin et al., 1997; Saunders et al., 2003). The impact of modifications in the chemical degradation mechanism suggested in a theoretical work by Vereecken et al. (2007) was tested. This includes the formation of new products and the change of branching ratios compared to the MCM.

1.1 Degradation mechanism for α -pinene

A simplified scheme giving the reactions most relevant in the experiments here is shown in Fig. 1. The α -pinene oxidation is initiated by the OH attack. As implemented in the MCM, OH adds to the carbon–carbon double bond of α -pinene, forming three different RO_2 radicals, APINAO $_2$, APINBO $_2$, and APINCO $_2$ (names and yields taken from the MCM).



According to the MCM, APINAO $_2$ and APINBO $_2$ are the mainly produced radicals of the reaction of α -pinene with OH with a contribution of 92 %, while APINCO $_2$ makes only a minor contribution of 8 %. Hydrogen abstraction by OH is not considered in the MCM. Consecutive reactions of the organic peroxy radicals with NO form alkoxy radicals, mostly APINAO and APINBO, which undergo a fast ring-opening and subsequent O_2 reaction yielding pinonaldehyde and HO_2 . This gives an overall pinonaldehyde yield of 84 % for these reaction pathways in the MCM under conditions of high NO. In contrast, the subsequent reaction of APINCO does not form pinonaldehyde, but product species include acetone and HCHO. Acetone and HCHO are also formed in the subsequent oxidation of pinonaldehyde, which is significant on the timescale of our experiments. Additional reactions of peroxy radicals forming nitrates and reactions with HO_2 are not shown.

Vereecken et al. (2007) investigated the reaction of α -pinene with OH using quantum-chemical calculations proposing modifications of branching ratios and additional reaction pathways. Firstly, three additional minor reaction channels of the attack of OH on α -pinene leading to an H abstraction are included. The total yield was calculated to be 12 %. These new pathways lead mainly to an increase in formaldehyde and also in acetone compared to the MCM. In contrast, less pinonaldehyde is formed. Secondly, the branching ratios of the other RO_2 species were revised. The addition of OH to the double bond was calculated to result in the

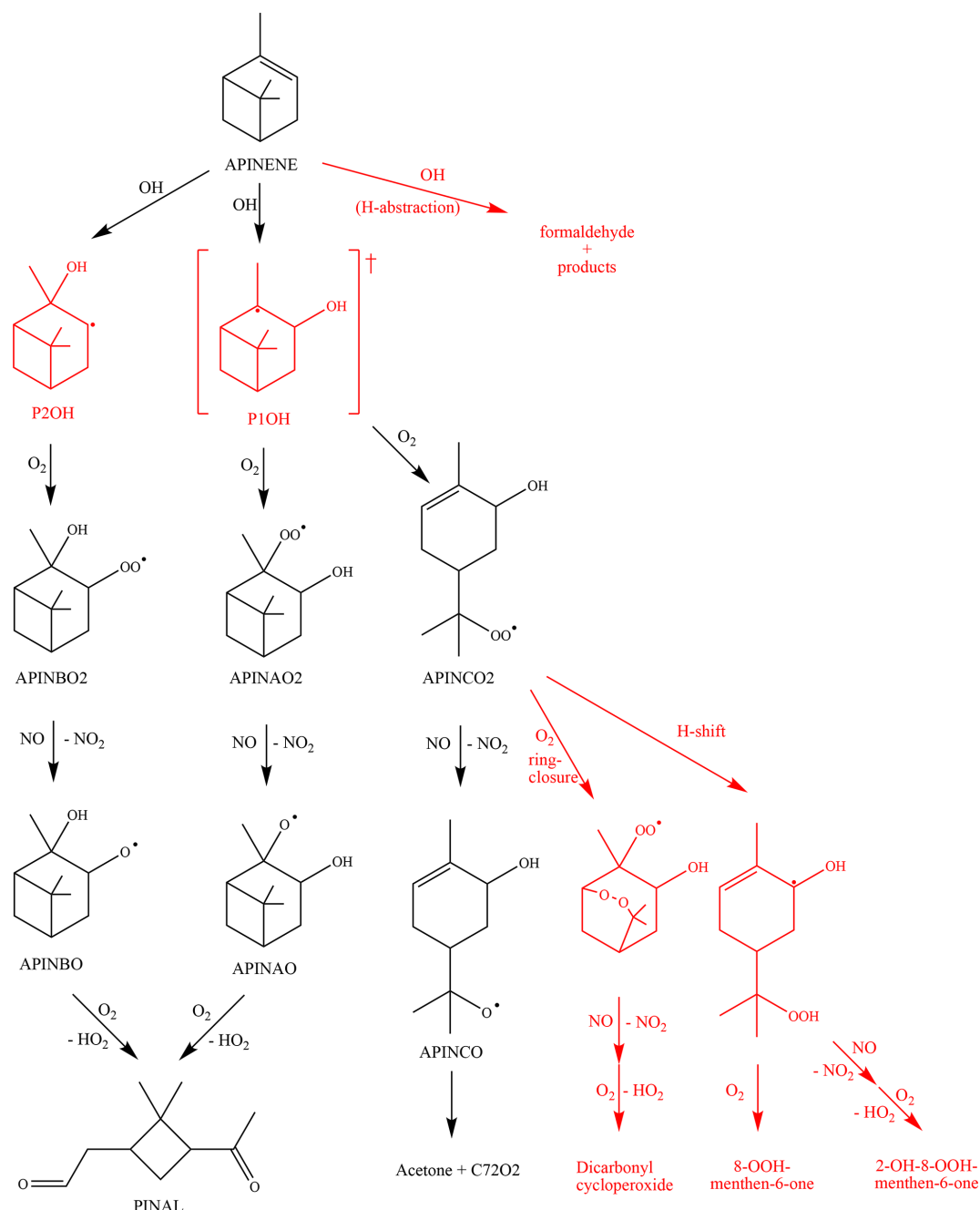


Figure 1. Simplified reaction of α -pinene as described in the MCM and modifications suggested by Vereecken et al. (2007) (shown in red). The hydrogen abstraction consists of three different pathways with a total contribution of 12 %. The ring-closure and H-shift reactions in the Vereecken mechanism outrun the formation of APINCO and therefore no acetone is directly formed in contrast to the MCM. RO₂ reactions with NO forming nitrate species are not shown. See text for details. The names are taken from the MCM (black) and according to Vereecken et al. (2007) (red).

OH attachment on both sites of attack, forming the adducts P1OH and P2OH with similar probability. P2OH further reacts with oxygen and forms the RO₂ radical APINBO2. Consequently, the APINBO2 yield is increased to 44 % compared to 35 % in the MCM. The tertiary radical P1OH is chemically activated. It is either thermally stabilized form-

ing the RO₂ radical APINAO2 after the O₂ addition or it undergoes a prompt ring-opening of the four member rings resulting in the RO₂ species APINCO2. The overall yield of APINAO2 and APINCO2 is 22 % each, suggested by Vereecken et al. (2007), compared to 57 % for APINAO2 and 7.5 % for APINCO2 assumed in the MCM. In addition,

Vereecken et al. (2007) calculated that a 1,6-H-shift reaction and a ring-closure reaction of APINCO₂ leading to 8-OOH-menthen-6-one, 2-OH-8-ooH-menthen-6-one, and a dicarbonyl cycloperoxide can compete with its reaction with NO. For conditions of the experiments of this work the dominant pathways are both unimolecular reactions, which are faster than the reaction with NO by a factor of at least 100. As a consequence, no acetone is directly formed in this pathway in contrast to the MCM.

Because the formation of APINCO₂ does not produce pinonaldehyde in the subsequent chemistry in contrast to APINAO₂ and APINBO₂, the shift in the RO₂ distribution in the mechanism by Vereecken et al. (2007) leads to an overall pinonaldehyde yield of 60 % compared to 84 % in the MCM.

2 Methods

2.1 Experiments in the simulation chamber SAPHIR

The experiments in this study were performed in the outdoor atmospheric simulation chamber SAPHIR at Forschungszentrum Jülich, Germany. The chamber has been described in detail before (e.g., Rohrer et al., 2005). The cylindrically shaped chamber (length 18 m, diameter 5 m, volume 270 m³) is made of a double-wall Teflon (FEP) film, which provides a high transmittance for the entire spectrum of solar radiation. A slight overpressure (30 Pa) in the chamber ensures that no air from the outside penetrates the chamber. A small flow of synthetic air that replaces the air sampled by the instruments and maintains the overpressure leads to a dilution of trace gases of approximately 4 % per hour. The synthetic air used for experiments is mixed from evaporated ultrapure liquid nitrogen and oxygen (Linde, purity ≥ 99.99990 %). Two fans inside the chamber are operated to ensure a rapid mixing of trace gases. A shutter system can keep the chamber dark, for example at the beginning of an experiment, and is opened to expose the chamber air to natural sunlight to perform photooxidation experiments. Small amounts of nitric acid (HONO), formaldehyde, and acetone are photolytically formed on the Teflon surface with source strengths of 100 to 200 pptv per hour when the chamber is illuminated by solar radiation (Rohrer et al., 2005). The primary source for OH radicals is the photolysis of HONO emitted by the chamber, also leading to a continuous increase in NO_x (= NO₂ + NO) in the experiment.

Two experiments were performed at similar conditions for this work: on 30 August 2012 at NO mixing ratios of less than 100 pptv and on 2 July 2014 at NO mixing ratios of less than 120 pptv. Before the experiments, the chamber was flushed with synthetic air until the concentrations of trace gases from previous experiments were below the detection limits of the instruments. The chamber air was humidified by flushing water vapor from boiling Milli-Q[®] water into the chamber together with a high flow of synthetic air. The rel-

ative humidity was approximately 70 % at the beginning of the experiments and decreased mainly due to the increase in the temperature during the day to approximately 20 % in the end of an experiment. A total of 40 ppbv ozone produced by a discharge ozonizer (O3Onia) was injected to simulate conditions typical for forested areas before the chamber roof was opened. In the first 2 h, the zero-air phase, no other reactive species were added, in order to quantify the small chamber sources for HONO, HCHO, acetone, and the background OH reactivity. Afterwards, α -pinene was injected from a high-concentration gas mixture of α -pinene in O₂ prepared in a SilcoNert-coated stainless steel canister (Restek) three times with time intervals of approximately 2 h. The time between the injections allowed us to study the photochemical degradation. The maximum α -pinene concentrations were 3.8 ppbv. After the initial phase, the OH reactivity was dominated by the injected α -pinene and its oxidation products, so that the background reactivity becomes secondary.

The additional loss of trace gases on the chamber surface is assumed to be negligible on the timescale of the experiment. The loss of α -pinene and pinonaldehyde, one of the oxidation products of α -pinene, was experimentally tested by injecting α -pinene and pinonaldehyde, respectively, into the clean chamber in the dark. The observed loss of these VOCs was consistent with the dilution due to the replenishment flow, demonstrating that there was no significant loss of α -pinene and pinonaldehyde on the Teflon film of the chamber.

2.2 Instrumentation

The set of instruments used in this work is listed in Table 1 giving the 1 σ accuracies and precisions.

OH was measured by laser-induced fluorescence (LIF) exciting OH at 308 nm (Holland et al., 1995; Fuchs et al., 2011). Previous studies reported interferences in the OH detection by LIF for some instruments (Mao et al., 2012; Novelli et al., 2014). A laboratory study investigating potential interferences from alkene ozonolysis reactions with the LIF instrument at SAPHIR (Fuchs et al., 2016) gave no hint for significant interferences for atmospherically relevant conditions. Only for exceptionally high, non-atmospheric reactant concentrations of ozone (300–900 ppbv) and some alkenes (1–450 ppbv) could interferences be observed. Hence, no interferences are expected for conditions of the experiments in this work. In addition, OH was detected by differential optical absorption spectroscopy (DOAS; Dorn et al., 1995). OH concentration measurements of both instruments agreed on average within 15 %. A similarly good agreement between both instruments has been found in previous studies (e.g., Schlosser et al., 2009; Fuchs et al., 2012).

The LIF instrument also measured the HO₂ concentrations in a second fluorescence cell, in which HO₂ is chemically converted to OH in a reaction with added NO. Fuchs et al. (2011) reported that this detection scheme can be affected by

interferences from organic peroxy radicals (RO_2) that also react with NO and rapidly form HO_2 . Consequently, in the experiments of this work, the NO concentrations were reduced to suppress the conversion of RO_2 as described in Fuchs et al. (2011) so that interferences become unimportant.

OH reactivity (k_{OH}), the inverse lifetime of OH, was measured with a pump-probe instrument (Lou et al., 2010; Fuchs et al., 2017). High OH concentrations are generated in a flow tube by laser flash photolysis of ozone in the presence of water and the decay of OH caused by ambient OH reactants is measured by LIF at the end of the flow-tube. The pseudo-first-order decay rate constant fitted to the time-resolved OH measurements gives the total OH reactivity. Unfortunately, OH reactivity could only be measured in the experiment conducted in 2012 because the instrument failed in the experiment in 2014.

α -pinene and oxygenated organic compounds expected to be formed in the α -pinene oxidation, acetone and pinonaldehyde, were detected by a proton-transfer-reaction time-of-flight mass spectrometer (PTR-TOF-MS; Lindinger et al., 1998; Jordan et al., 2009). However, only for one of the experiments (2 July, 2014) was the PTR-TOF-MS calibrated to quantify pinonaldehyde. In addition, a gas chromatograph with a flame ionization detector (GC-FID) was used for the measurements of α -pinene and acetone. VOC concentrations measured by a GC-FID were on average 25 % lower than those measured by a PTR-TOF-MS for the experiment conducted in 2012 and 15 % lower for the experiment in 2014. This discrepancy needs to be taken into account as additional uncertainty. Formaldehyde that is also expected to be produced in the oxidation of α -pinene was measured by a Hantzsch monitor and by differential optical absorption spectroscopy. The measured concentrations agreed on average within 6 %.

CO and water vapor mixing ratios were monitored with a cavity ring-down instrument (Picarro), NO and NO_2 using a chemiluminescence instrument (Eco Physics), and O_3 with an UV absorption instrument (Ansyco). Photolysis frequencies were calculated from solar actinic flux densities measured by a spectroradiometer (Bohn et al., 2005; Bohn and Zilken, 2005).

2.3 Model calculations

The time series of trace-gas compounds and radicals were calculated by a zero-dimensional box model applying the chemistry of the Master Chemical Mechanism in the recent version 3.3.1.

The MCM was extended by chamber-specific processes like dilution and small sources of HONO, formaldehyde, acetaldehyde, and acetone that are present in the sunlit chamber (Rohrer et al., 2005). The chamber sources were implemented as continuous sources that are parameterized by temperature, relative humidity, and radiation as described by Rohrer et al. (2005). This function was scaled for each exper-

iment, such that the observed acetone and HCHO time series were matched during the zero air phase when no chemical production was expected. It was assumed that the scaling factors remained constant over the course of one experiment.

The dilution rate was calculated from the monitored replenishment flow rate. NO, NO_2 , HONO, water vapor mixing ratio, temperature, and pressure were constrained to measurements. While photolysis frequencies for NO_2 , HONO, O_3 , and pinonaldehyde were calculated from actinic flux measurement, all other photolysis frequencies were calculated for clear-sky conditions as parameterized in MCM 3.3.1 but scaled by the ratio of measured to calculated $j(\text{NO}_2)$ to account for cloud coverage and chamber effects. The pinonaldehyde photolysis frequency was calculated, based on the measured absorption cross sections of pinonaldehyde by Hallquist et al. (1997) and a quantum yield of 1. These photolysis frequencies are greater by a factor of 3.5 compared to the parameterization in the MCM.

Modeled parameters were calculated on a 1 min time base. α -pinene and O_3 injections were introduced as sources only present at the time of injection. The O_3 source strengths were adjusted to match measurements at the time of the injection. Similarly, the α -pinene source was adjusted to the increase in the OH reactivity. For the experiment where no k_{OH} measurements were available, the increased α -pinene concentrations measured by a PTR-TOF-MS were used instead.

Sensitivity studies (M1) were performed applying modifications of the MCM based on a theoretical study by Vereecken et al. (2007). An overview of the model modifications applied to the MCM is given in Tables S2 and S3 in the Supplement. The mechanism based on Vereecken et al. (2007) differs from the MCM in new pathways, branching ratios, and product yields. RO_2 radicals formed from the reactions of this modified mechanism are assumed to react similar to other RO_2 species with NO, HO_2 , and other RO_2 species. Additional first-generation oxygenated organic compounds are formed that are not part of the MCM. In the model run denoted M1, no further reactions of these products were implemented.

Branching ratios of the initial OH + α -pinene reaction were further adjusted in model run M2 to better match observations. An overview of the simplified reaction scheme indicating the differences between the model runs is shown in Table 3.

3 Results and discussion

3.1 Product yields

The fate of the RO_2 and therefore also product yields depends on the NO concentration. At low NO levels, the reaction of RO_2 with HO_2 and RO_2 recombination reactions can compete with the reaction of RO_2 with NO. For NO mixing ratios of up to 120 ppt in the experiment on 2 July 2014,

approximately 70 % of the RO₂ radicals reacted with NO and 30 % with HO₂, while RO₂ self-reactions were not significant. Product species quantified in the experiments were mainly formed in the reaction of RO₂ with NO. In contrast, the reaction of RO₂ with HO₂ terminated the radical chain reactions and forms hydroperoxide species (ROOH) that were not detected.

Product yields were calculated from measured product concentrations in relation to the α -pinene that reacted with OH. The α -pinene and pinonaldehyde concentrations were determined by a PTR-TOF-MS. Acetone concentrations were derived from interpolated GC-FID data to exclude possible interferences on the quantifier ion of acetone in the PTR-TOF-MS. Because products were further oxidized in the experiment or partly had sources not related to the α -pinene chemistry, a correction was applied. The correction procedure used here follows the description in Galloway et al. (2011) and Kaminski et al. (2017). The α -pinene reacted away was corrected for dilution and its reaction with O₃. Ozonolysis accounted for approximately 25 % of the loss of α -pinene. Product concentrations were corrected for their loss from photolysis, from their reaction with OH, and from dilution. In addition, their production from the chamber sources and from the α -pinene ozonolysis was subtracted from the measured concentrations. Acetone and HCHO chamber source strengths were determined in the initial phase of each experiment when the chamber air was already exposed to sunlight, but before the injection of α -pinene. The production rates were 0.04 and between 0.11 and 0.27 ppbv h⁻¹ for acetone and formaldehyde, respectively. For acetone the chamber source contributed only 10 % to the overall formed acetone. In contrast, up to 60 % of the total measured HCHO was formed on the chamber walls, which could lead to an additional bias of the determined yield. A detailed description of the corrections is given in the Supplement. Yields and reaction rate constants used for the correction were taken from recommendations in Atkinson et al. (2006) also used in the MCM.

Figure 2 shows the relation between the consumed α -pinene and product concentrations. The product yields were determined from the slopes of the relation resulting in yields for pinonaldehyde of (5 ± 3) %, for acetone of (19 ± 6) %, and for formaldehyde of (11 ± 5) %. The uncertainty for the pinonaldehyde yield is 1 σ derived from measurements and errors of the applied correction in one experiment in 2014, when pinonaldehyde was quantified. The stated acetone and HCHO yields are the combined result from both experiments in 2012 and 2014 and the error gives the range of values derived in the two experiments. The relationships between consumed α -pinene and acetone and formaldehyde are not exactly linear because both are not only directly formed in the reaction of α -pinene with OH, but also from the subsequent oxidation of products such as pinonaldehyde. Therefore, acetone and formaldehyde yields increase over the course of the experiment. The slope at the early stage of the experi-

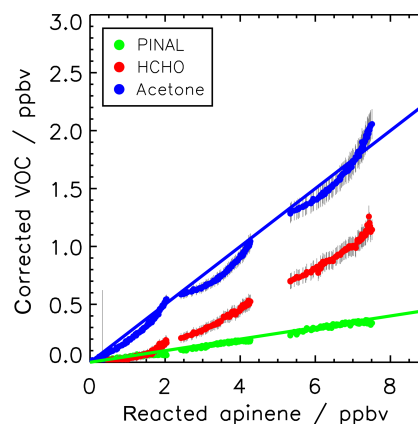


Figure 2. Yield of pinonaldehyde, acetone, and formaldehyde determined from the slope of the relation between consumed α -pinene and measured oxidation product concentrations for the experiment on 2 July 2014. Organic oxidation product concentrations are corrected for losses and production not related to the α -pinene + OH reaction (see text for details). Colored lines give the results of a linear regression. The HCHO yield is determined from the initial slope but increases at later times of the experiments as indicated by the increasing slope of the relationship.

ment, when only a little α -pinene reacted away, best reflects their formation yield directly from the α -pinene + OH reaction, whereas the slope at later times gives the overall yield of the α -pinene degradation. The nonlinear behavior is most strongly pronounced for formaldehyde, for which the yield increased from 5 % to 20 % over the course of the experiment.

Results of this work are compared to results from previous studies in Table 2. Yields of pinonaldehyde that were detected by various instruments in previous studies were highly variable ranging from 6 % to 87 %. High yields of 37 %–87 % are reported in the studies by Hatakeyama et al. (1991) and Nozière et al. (1999), in which Fourier-transform infrared spectroscopy (FT-IR) was applied. Larsen et al. (2001) found a pinonaldehyde yield of 6 % also measured using FT-IR. FT-IR measurements may suffer from interferences from other carbonyl compounds, which could have led to overestimated yields (Eddingsaas et al., 2012). Studies measuring pinonaldehyde with a GC-FID (Arey et al., 1990; Hakola et al., 1994; Jaoui and Kamens, 2001; Aschmann et al., 2002; Lee et al., 2006) and PTR-MS (Lee et al., 2006; Wisthaler et al., 2001) gave similar yields that are in the range of 28 % to 34 %.

Except for some experiments performed in the absence of NO, experiments in previous studies were carried out under conditions when RO₂ reacted mainly with NO.

In a recent study by Isaacman-VanWertz et al. (2018), the carbon budget was analyzed during the photooxidation of α -pinene by OH making use of various mass spectrometry instruments. While the carbon budget was found to be closed

Table 1. Instrumentation for radical and trace-gas measurements.

Species	Technique	Time resolution	1 σ precision	1 σ accuracy
OH	DOAS ^a	205 s	$0.6 \times 10^6 \text{ cm}^{-3}$	6.5 %
OH	LIF ^b	47 s	$0.6 \times 10^6 \text{ cm}^{-3}$	13 %
HO ₂	LIF ^b	47 s	$1.5 \times 10^7 \text{ cm}^{-3}$	16 %
<i>k</i> _{OH}	Laser photolysis + LIF ^b	180 s	0.3 s^{-1}	0.5 s^{-1}
NO	Chemiluminescence	180 s	4 pptv	5 %
NO ₂	Chemiluminescence	180 s	2 pptv	5 %
O ₃	UV absorption	10 s	1 pptv	5 %
α -pinene, pinonaldehyde, and acetone	PTR-TOF-MS ^c	40 s	15 pptv	14 %
α -pinene and acetone	GC-FID ^d	30 min	(4–8) %	5 %
HONO	LOPAP ^e	300 s	1.3 pptv	10 %
HCHO	DOAS ^a	100 s	20 %	10 %
Photolysis freq.	Spectroradiometer	60 s	10 %	10 %

^a DOAS: differential optical absorption spectroscopy. ^b LIFL: laser-induced fluorescence. ^c PTR-TOF-MS: proton-transfer-reaction time-of-flight mass spectrometer. ^d GC-FID: gas chromatograph–flame ionization detector. ^e LOPAP: long-path-absorption photometer.

at the end of their experiment, the initial phase showed a discrepancy of up to 30 % between measured species including pinonaldehyde and α -pinene that reacted away. This indicates that a substantial fraction of products was not detected, consistent with low pinonaldehyde yields found in this and previous studies. Although no pinonaldehyde yield was reported, the yield can be estimated to be less than 20 % from figures shown in Isaacman-VanWertz et al. (2018).

The pinonaldehyde yield in the study here agrees within the stated errors with the yields reported by Larsen et al. (2001), but is lower than in all other previous studies, which used significantly higher α -pinene mixing ratios of hundreds of parts per billion by volume or even several parts per million by volume. Concentrations of this work were close to those typically found in ambient air. This appears to be the major difference between the experiments here and previous experiments.

The acetone yield in this study of 19 ± 6 % is higher by nearly a factor of 2 compared to previously reported values. The higher acetone yield corresponds to the lower pinonaldehyde yield and could therefore be a result of reaction pathways that do not lead to the formation of pinonaldehyde but forming acetone instead.

Only a few of the previous studies reported formaldehyde yields. The yield determined in this study agrees within the stated uncertainty with values in Nozière et al. (1999) (zero NO), Larsen et al. (2001), Wisthaler et al. (2001), and Lee et al. (2006), but is significantly lower than the yield in the studies by Hatakeyama et al. (1991) and Nozière et al. (1999) (high NO). Like for acetone, additional pathways not included in the mechanism could lead to HCHO formation instead of pinonaldehyde. It is also not clear if HCHO yields of the different studies are comparable because the yield increases, whether organic products are further oxidized during the experiment (Fig. 2). The studies by Hatakeyama et al. (1991) and Nozière et al. (1999) (high NO) were performed

at high reactant concentrations and high NO concentrations, which accelerate the oxidation rate.

Few studies were performed in the presence of water, which can have an impact on product yields as shown for the product yields in the ozonolysis of α -pinene (Tillmann et al., 2010). A water dependence in the OH degradation mechanism of α -pinene has not been reported yet. In general, yields of products strongly depend on the fate of RO₂. Most of the studies were performed at high reactant concentrations and also in the presence of high NO concentrations. Therefore RO₂ recombination reactions might have played a larger role compared to the chamber experiment here. In addition, the fast oxidation of α -pinene led to particle formation in some of the studies, and therefore additional heterogeneous chemistry affected the results (e.g., Nozière et al., 1999). The chamber study here was performed at atmospheric reactant concentrations such that the RO₂ lifetime was approximately 0.5 min with respect to reactions with both HO₂ and NO, but was long enough that potential isomerization reactions could compete. The differences in the RO₂ fate likely explain the large variety of yields in the different studies. This demonstrates the importance of performing experiments at atmospheric levels of reactants as done in this study.

3.2 Comparison of trace-gas measurements with MCM 3.3.1 model calculations

Time series of measured species are compared to model calculations using the MCM for the experiment conducted on 2 July 2014 (Figs. 3 and 4). This experiment is discussed here in more detail because measurements of pinonaldehyde were available in contrast to the experiment in 2012. Time series for the experiment in 2012 are shown in Fig. 5.

After the first α -pinene injection, OH chemistry is dominated by reactions with α -pinene. Thereby formed pinonaldehyde is overestimated in the MCM by a factor of 4.

Table 2. Yields of pinonaldehyde, formaldehyde, and acetone for the reaction of α -pinene + OH compared to literature values. Experimental conditions and applied measurement technique for the detection of organic compounds are additionally listed.

Reference	Yield (%)			Exp. conditions			
	Pinonaldehyde	Acetone	HCHO	α -pinene (ppbv)	NO (ppbv)	Water (rH %)	Technique
Arey et al. (1990)	29	–	–	400–900	10 000	0	GC–FID
Hatakeyama et al. (1991)	56 $\pm$ 4	–	54 $\pm$ 5	950–1300	390–2300	9	FT-IR
Hakola et al. (1994)	28 $\pm$ 5	–	–	350–1000	10 000	0	GC–FID
Nozière et al. (1999)	87 $\pm$ 20	9 $\pm$ 6	23 $\pm$ 9	200–2700	4000	0	FT-IR
Jaoui and Kamens (2001)	28	–	–	940–980	430–490	18–40	Denuder, GC–MS
Larsen et al. (2001)	6 $\pm$ 2	11 $\pm$ 3	8 $\pm$ 1	1400–1600	1000	2–5	FT-IR
Aschmann et al. (2002)	28 $\pm$ 5	–	–	400–900	7000–9000	0	GC–FID
Lee et al. (2006)	30 $\pm$ 0.3	6	16	109	9	0	PTR-MS
Nozière et al. (1999)	37 $\pm$ 7	7 $\pm$ 2	8 $\pm$ 1	200–2700	NO free	0	FT-IR
Wisthaler et al. (2001)	34 $\pm$ 9	11 $\pm$ 2	8 $\pm$ 1	1000–1300	NO free	0	PTR-MS
this work	5 $\pm$ 3 ^a	19 $\pm$ 6 ^b	11 $\pm$ 5 ^b	3.8	< 0.1	30–60	PTR-TOF-MS

^a Yield determined in the 2014 experiment. ^b Combined yield from experiments in 2012 and 2014.

Table 3. RO₂ yields for the different model runs and the resulting pinonaldehyde yields for the chamber experiments.

Model run	Yield (%)				
	APINAO2	APINBO2	APINCO2	Pinonaldehyde	
MCM	57.2	35.3	7.5	84	MCM 3.1.1
M1	22	44	22	60	Vereecken et al. (2007)
M2	0	5	83	5	adjusted to the measured pinonaldehyde yield

The pinonaldehyde concentrations increase directly after the VOC injections but start to decrease 1 h later due to its consumption by photolysis and reaction with OH. The model underestimates OH and HO₂ concentrations by approximately 25 %. Because too much pinonaldehyde is formed in the model, the OH consumption is overestimated, which can partly be the reason for the smaller OH concentrations than observed.

Three α -pinene injections with concentrations of 2–3 ppbv each were performed. The modeled α -pinene consumption is slightly slower by approximately 10 % than measured, consistent with the lower modeled OH compared to measured OH. This is also seen in a slower decrease in the modeled OH reactivity compared to measurements carried out in the experiment in 2012 (Fig. 5). The fact that OH reactivity is dominated by α -pinene specifically shortly after its injection supports that α -pinene decays are slower in the model compared to observations.

The production of acetone in the model matches the observations within the stated errors. In contrast, the formation of formaldehyde is slightly overestimated by around 10 %.

Modeled and measured O₃ concentrations start to slightly deviate in the second half of the experiment but agree over the whole experiment within the measurement uncertainty.

3.3 Sensitivity model calculations

Sensitivity model runs were performed to test if shortcomings of the MCM model results can be explained by either recent studies reported in literature or further adjustments.

Figure 3 shows in orange a sensitivity run with HO₂ and pinonaldehyde concentrations constrained to measurements. Modeled and measured OH concentrations agree within the stated uncertainty and the time behavior is reproduced in contrast to the MCM model run. This indicates that the radical budget is closed. As a result of the higher OH concentration the α -pinene is consumed faster compared to the MCM and the resulting decay reproduces the observations within the measurement uncertainty. Constraining only the HO₂ data is not sufficient because the OH loss by pinonaldehyde would be overestimated.

The application of the mechanism by Vereecken et al. (2007) reduces the pinonaldehyde yield by 24 % compared to the MCM, reducing the model–measurement discrepancy by a factor of 2 (Fig. 4). The lower pinonaldehyde yield also results in an increased OH concentration because decomposition products like acetone, which reacts more slowly with OH, are produced instead. Therefore, the OH loss rate constant is reduced compared to the results obtained with the MCM. This reduces the model–measurement discrepancy to 10 %. As discussed above, higher OH concentrations also

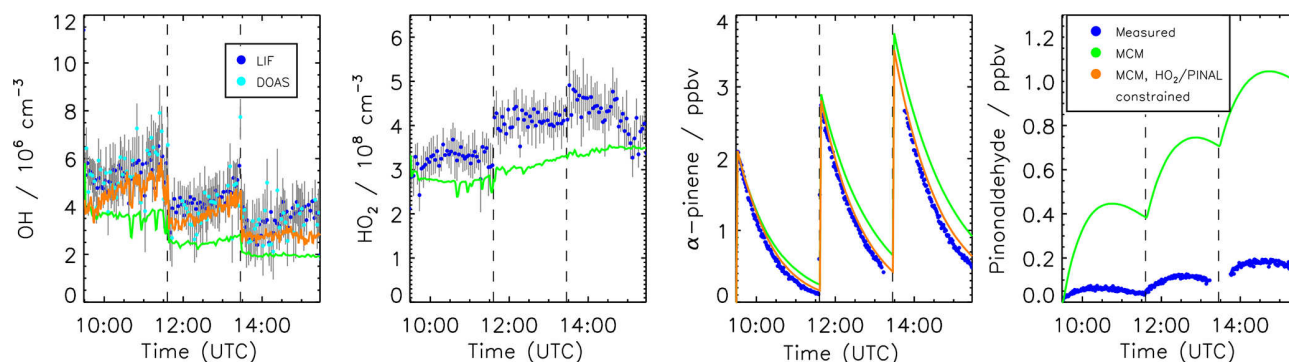


Figure 3. Time series of measured and modeled concentrations of radicals, α -pinene, and pinonaldehyde for an MCM model run with and without having HO_2 and pinonaldehyde constrained to measurements (experiment on 2 July 2014).

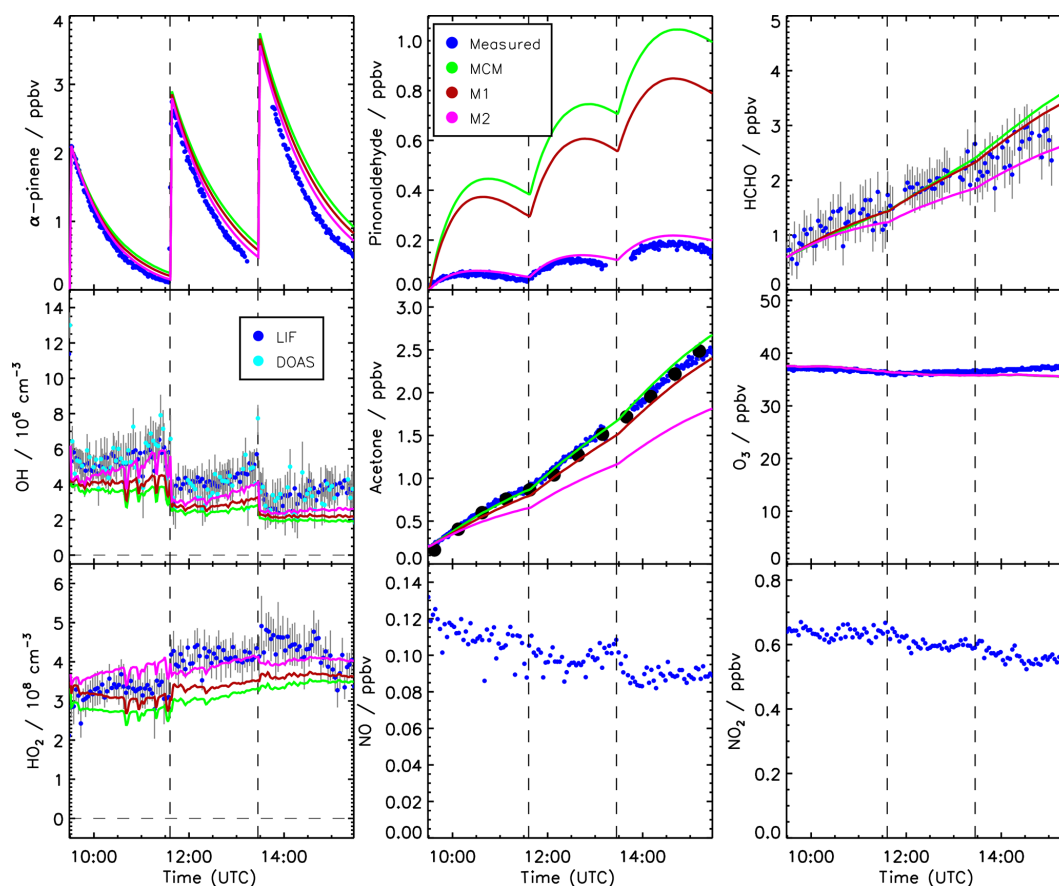


Figure 4. Time series of measured and modeled concentrations of radicals and inorganic and organic compounds during the α -pinene photooxidation at low NO (experiment on 2 July 2014).

lead to a faster consumption of the α -pinene, so that the model–measurement agreement of α -pinene is also improved to within 5 %. This is also consistent with results obtained in the experiment performed in 2012, when OH reactivity was measured. OH reactivity is approximately 20 % higher than predicted by the MCM at the end of the experiment, whereas agreement within 8 % is achieved when modifications by

Vereecken et al. (2007) are applied. However, there is more uncertainty in the total OH reactivity determined from the model because reaction rate constants of the products dicarbonyl cycloperoxide, 8-OOH-menthen-6-one, and 2-OH-8-OOH-menthen-6-one that are formed instead of pinonaldehyde from the subsequent reaction of APINCO₂ can only be

estimated. Here, a similar reaction rate constant like the one for pinonaldehyde is assumed.

The modifications suggested by Vereecken et al. (2007) reduce the model–measurement discrepancies for radicals, OH reactivity, α -pinene, and pinonaldehyde without changing the reasonable agreement for formaldehyde and acetone, but measured pinonaldehyde concentrations are still significantly lower than predicted by the model. Because APINCO₂ is the only RO₂ species that does not form pinonaldehyde, another sensitivity study (M2) was performed for which the RO₂ distribution is adjusted, such that modeled pinonaldehyde concentrations match observations. This requires a yield of APINCO₂ of 86 % making the prompt ring-opening reaction subsequent to the OH attachment to α -pinene the most important pathway. The yields for the other two RO₂ species are consequently reduced to 0 % and 5 % for APINAO₂ and APINBO₂, respectively. The minor reaction pathways suggested by Vereecken et al. (2007) (H abstractions) remained unchanged in this model run. A similar shift in the RO₂ distribution towards APINCO₂ was proposed by Xu et al. (2019). The authors reported a branching ratio of 69 % for the initial OH addition forming P1OH and a branching ratio of 97 % for the subsequent ring-opening reaction. The resulting overall APINCO₂ yield was 60 % (see Supplement).

If this model modification is applied, HO₂ radical concentrations are increased by up to 30 % giving reasonable agreement within the stated uncertainties between modeled values and measurements (Fig. 4). The increased HO₂ together with a reduced OH loss rate due to the decreased pinonaldehyde concentration result in up to 30 % higher modeled OH radical concentrations compared to M1, which agree with measurements within the measurement uncertainty. In M2, acetone and formaldehyde are now underestimated by approximately 20 % and 10 %, respectively, because the production by the photooxidation of pinonaldehyde is decreased. The underestimation of acetone and formaldehyde may be caused by the missing unknown degradation chemistry of products which are postulated in the mechanism by Vereecken et al. (2007). Vereecken et al. (2007) suggest that, for example, dicarbonyl cycloperoxide formed in the revised oxidation scheme likely produces acetone.

The change of the RO₂ yields is only one possibility to match the measured pinonaldehyde mixing ratios and does not imply that this is the correct oxidation scheme. If the initial RO₂ branching ratio suggested by Vereecken et al. (2007) is correct, then unknown reactions of APINAO₂ and APINBO₂ or APINAO and APINBO, which suppress the pinonaldehyde formation, could also explain the discrepancies. These unknown reactions need to be significantly faster than the currently known reactions to compete with the other reactants NO, HO₂, and RO₂.

3.4 Comparison with previous studies

The oxidation scheme of β -pinene that has a similar structure as α -pinene has previously been investigated in the SAPHIR chamber for comparable conditions (Kaminski et al., 2017) giving similar results obtained for α -pinene here. α - and β -pinene are isomers, which differ from each other by the position of the double bond. The double bond is endocyclic in α -pinene and exocyclic in β -pinene. Like the assumed major oxidation product, pinonaldehyde, for α -pinene, the main oxidation product of β -pinene, nopinone, was found to be overestimated by up to a factor of 3 using the MCM model calculations in Kaminski et al. (2017). Similar to α -pinene, Vereecken and Peeters (2012) suggested a dominant ring-opening reaction after the addition of OH to the double bond of β -pinene that leads to products other than nopinone. The chamber study by Kaminski et al. (2017) confirms that the measured nopinone concentration is consistent with this mechanism. In addition, these reaction pathways can lead to a faster production of HO₂ and improve the model–measurement agreement of OH and HO₂ concentrations.

Two field campaigns in environments in which monoterpene species were the dominant reactive organic compounds showed large discrepancies between modeled and observed OH and HO₂ radical concentrations. During the Bio-hydro-atmosphere interactions of Energy, Aerosols, Carbon, H₂O, Organics, and Nitrogen-Rocky Mountain Organic Carbon Study (BEACHON-ROCS) campaign in 2010 in a forested area, the main biogenic organic compounds were 2-methyl-3-butene-2-ol (MBO) and monoterpenes with average mixing ratios of 1.6 ppbv and 0.5 ppbv, respectively. Model calculations conducted with the University of Washington Chemical Model (UWCM) underestimated HO₂ radical concentrations by up to a factor of 3 and OH concentrations could only be reproduced by the model, if HO₂ was constrained to measurements (Kim et al., 2013).

During the HUMPPA-COPEC (Hyytiälä United Measurements of Photochemistry and Particles in Air – Comprehensive Organic Precursor Emission and Concentration study) field campaign in 2010 in a boreal forest in Finland, α -pinene mixing ratios peaked around 1 ppbv. Again, the model calculations with CAABA/MECCA (Chemistry As A Boxmodel Application/Module Efficiently Calculating the Chemistry of the Atmosphere) gave similar results as reported by Kim et al. (2013) for the site in the Rocky Mountains. Modeled k_{OH} and HO₂ concentrations were both underestimated by model calculations. Hens et al. (2014) attributed this to a missing HO₂ source.

Results from both field campaigns are consistent with findings in the chamber experiments for α -pinene here and β -pinene reported by Kaminski et al. (2017). MBO was also an important OH reactant during the BEACHON-ROCS campaign, but is likely not responsible for the observed model–measurement discrepancies. A chamber experiment reported

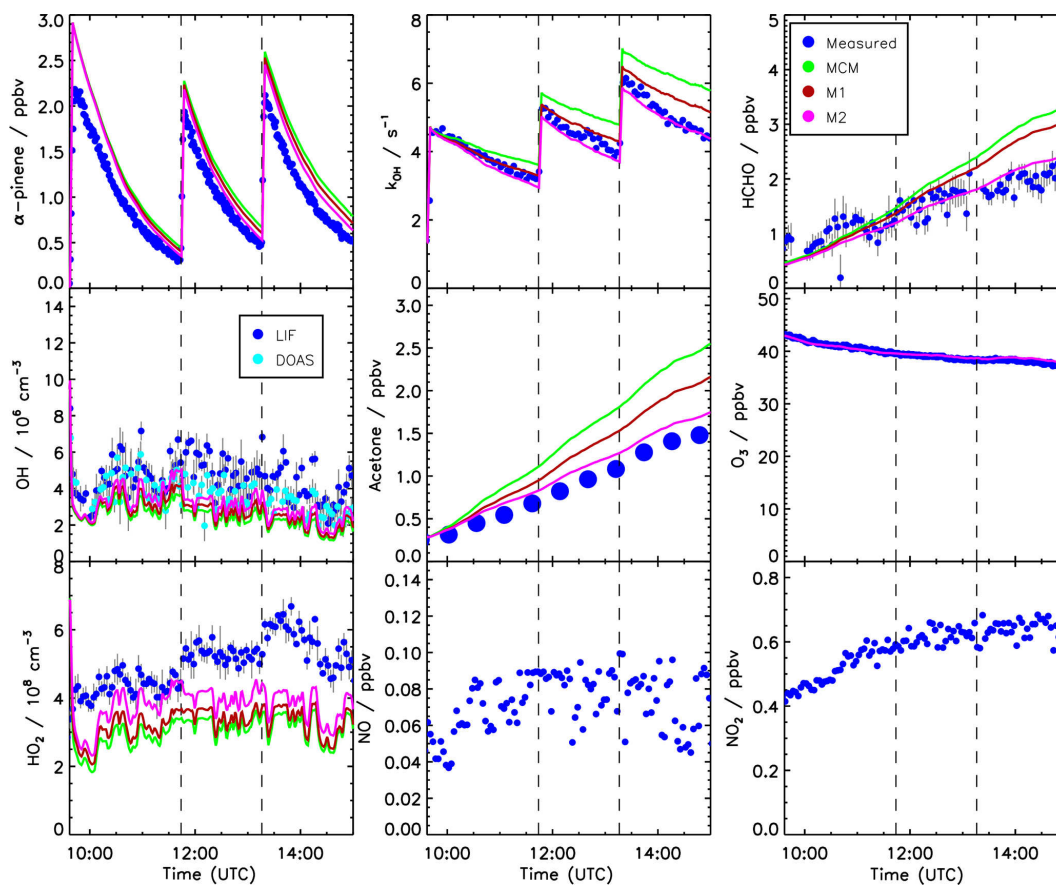


Figure 5. Time series of measured and modeled concentrations of radicals and inorganic and organic compounds during the α -pinene photooxidation at low NO (experiment on 30 August 2012).

by Novelli et al. (2018) demonstrated that the radical budget in the oxidation of MBO is well understood by current chemical models. In addition, quantum-chemical calculations by Knap et al. (2016) showed that H-shift isomerization reactions are negligible for RO_2 radicals formed in the reaction of OH with MBO.

In a recent laboratory study, Eddingsaas et al. (2012) investigated the α -pinene oxidation by OH at low NO_x conditions, when the fate of RO_2 radicals was dominated by $\text{RO}_2 + \text{HO}_2$ reactions. The authors suggest that pinonaldehyde is formed from $\text{RO}_2 + \text{HO}_2$ through an alkoxy radical channel that regenerates OH (Eddingsaas et al., 2012) and from the photooxidation and photolysis of α -pinene hydroxy hydroperoxides formed in this reaction. Pinonaldehyde yields were not measured but estimated to be around 33 % for low NO_x . These results do not contradict results here because only approximately 30 % of the RO_2 reacted with HO_2 in the chamber experiment and less hydroxy hydroperoxides are formed.

Recently Xu et al. (2019) evaluated unimolecular reaction pathways in the photooxidation of α -pinene by OH. The authors described that the hydroxy group and carbon–carbon double bond found in APINCO2 enhances the rates of uni-

molecular reactions. This is consistent with the faster HO_2 production observed in this work that was reproduced by the model after increasing the branching ratio of the APINCO2 formation. Xu et al. (2019) further reported that unimolecular reactions in the α -pinene degradation do not convert NO to NO_2 and therefore impact the O_3 formation. For our experiment conditions no discrepancies in the O_3 concentrations are observed.

4 Summary and conclusions

The photooxidation of α -pinene was investigated in the atmospheric simulation chamber SAPHIR. Two experiments were performed under atmospheric α -pinene concentrations (≤ 3.8 ppbv) and medium NO conditions (≤ 120 pptv). Measured time series were compared to model results applying the recent version of the Master Chemical Mechanism version 3.3.1.

Model calculations lead to approximately 25 % lower OH and HO_2 radical concentrations than measured. In addition, pinonaldehyde is the major organic oxidation product in the MCM (84 %), whereas the measured pinonaldehyde

yield is only $(5 \pm 3) \%$ in the chamber experiment. This is in the lower range of previous pinonaldehyde yields determined in laboratory experiments which range between 6 % and 87 %. This large range might reflect the variety of conditions in the experiments. In addition, laboratory studies were often performed at high NO and α -pinene concentrations. The chamber study in this work is the first one using atmospheric conditions of reactant concentrations. Yields of acetone (0.19 ± 0.06) and formaldehyde (0.11 ± 0.05) in this study are reproduced by model calculations applying the MCM.

Reaction pathways from quantum-chemical calculations by Vereecken et al. (2007) were implemented in sensitivity model runs leading to a reduction in the model–measurement discrepancies for the radical and pinonaldehyde concentrations by approximately 10 % and 25 %, respectively. The major change is due to a shift in the branching ratio of the RO₂ distribution after the OH addition to α -pinene, favoring reaction pathways that do not lead to the production of pinonaldehyde. Further adjustments of this distribution can bring model predictions into agreement with observations, but other unknown reaction pathways that reduce the pinonaldehyde yield could also explain the observations.

A chamber study on β -pinene by Kaminski et al. (2017), supported by quantum-chemical calculations by Vereecken and Peeters (2012), reported that a similar shift in the initial RO₂ branching ratio towards a ring-opening reaction was needed to explain product distribution and radical concentrations. Results are consistent with findings in field studies (Kim et al., 2013; Hens et al., 2014) where monoterpene emissions were high. Similar to the chamber studies for α - and β -pinene, models gave significantly less OH and HO₂ compared to measured values.

Further experiments with a more detailed analysis of the organic oxidation products could help to clarify the exact reaction mechanism and further support results from quantum-chemical calculations.

Data availability. Data of the experiments in the SAPHIR chamber used in this work are available on the EUROCHAMP data home page (<https://data.eurochamp.org/>, last access: 28 April 2019, Eurochamp, 2019).

Supplement. The supplement related to this article is available online at: <https://doi.org/10.5194/acp-19-11635-2019-supplement>.

Author contributions. MR analyzed the data and wrote the paper. HF and MK designed the experiments. HF conducted the HO_x radical measurements and SN was responsible for the OH reactivity measurements. BB conducted the radiation measurements. MK and RW were responsible for the GC measurements. RT, AL, and IHA were responsible for the PTR-TOF-MS measurements. XL was responsible for the HONO measurements and HPD for the DOAS OH

data. FR was responsible for the NO_x and O₃ data. All co-authors commented on the paper.

Competing interests. The authors declare that they have no conflict of interest.

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